

## Chemistry 2<sup>nd</sup> Semester Final Review **KEY**

\*This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.\*

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

*Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)*

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)  
 $V_1=3.8 \text{ L}; P_1=2.71 \text{ atm}$   
 $V_2=1.47 \text{ L}; P_2=? \text{ Atm}$   
Since all units are standard (L, atm), we do not need to convert the units before proceeding.  
 $(P_1V_1)/T_1=(P_2V_2)/T_2$   
T goes away (not mentioned in problem)  
 $P_1V_1=P_2V_2$   
Isolate for P<sub>2</sub>:  $P_2=(P_1V_1)/V_2$   
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})=7.0 \text{ atm}$  (L cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)  
 $V_1=1.49 \text{ L}, T_1=34.75 \text{ degrees Celsius}$   
 $V_2=? \text{ L}, T_2=78.41 \text{ degrees Celsius}$   
We need to convert from Celsius to Kelvins before proceeding.  
 $T_1=34.75 + 273=307.75 \text{ K}$   
 $T_2=78.41 + 273=351.41 \text{ K}$   
 $(P_1V_1)/T_1=(P_2V_2)/T_2$   
Get rid of P  
 $V_1/T_1=V_2/T_2$   
Isolate V<sub>2</sub>:  $V_2=(V_1/T_1)T_2$   
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})=1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?  
The combined gas law won't work here. You only have one condition and nothing changes.  
 $PV=nRT$  will work. (n= moles and R= 0.08206 L atm/K mol)  
 $P=1.04 \text{ atm}, V=? \text{ L}, n=?$  OOPS we need to convert from grams to moles!,  $T=84.7 \text{ degrees Celsius} + 273=357.7 \text{ K}$   
n: Hydrogen gas is diatomic!  
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/2.02 \text{ g H}_2)=4.193069307 \text{ mol H}_2$   
Isolate V:  $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K})/1.04 \text{ atm}=118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

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