

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa from [OH<sup>-</sup>] and [H<sup>+</sup>].
3. Calculate the pH of weak (or strong) acids.

## Vocabulary

- |                   |   |                   |
|-------------------|---|-------------------|
| 1. autoionization | 2. conjugate acid-base pair                       | 3. acid solution  |
| 4. hydroxide ion  | 5. hydronium ion (H <sub>3</sub> O <sup>+</sup> ) | 6. autoionization |
|                   | 7. acid solution                                  | 8. pH             |

## Key Equations

- |  |                        |
|--|------------------------|
| 1. $[H^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H^+]$   |
| 3. $pH = -\log[H^+]$                             | 4. $pOH = -\log[OH^-]$ |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of one. Complete as needed.

1. What product is the \_\_\_\_\_ of a weak hydroxide ion and a hydroxide ion? \_\_\_\_\_ of the conjugate acid-base pair.
2. \_\_\_\_\_ of a strong acid is \_\_\_\_\_.
3. \_\_\_\_\_ of a weak acid is \_\_\_\_\_.
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# 19 2 Hydrogen Ions And Acidity Section Review Answers

**K Morrison**



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