

CHEMISTRY

Formulas

TEMPERATURE CONVERSIONS

$$F^{\circ} \rightarrow C^{\circ}$$

$$C^{\circ} = \frac{5}{9}(F^{\circ} - 32)$$

$$C^{\circ} \rightarrow F^{\circ}$$

$$F^{\circ} = \frac{9}{5}C^{\circ} + 32$$

$$C^{\circ} \rightarrow K^{\circ}$$

$$K^{\circ} = C^{\circ} + 273.15$$

$$K^{\circ} \rightarrow C^{\circ}$$

$$C^{\circ} = K^{\circ} - 273.15$$

AVOGADRO'S NUMBER

$$1 \text{ mole} = 6.02 \times 10^{23}$$

PERCENT COMPOSITION

$$\text{percent composition} = \frac{\text{total molar mass of just the element}}{\text{molar mass of the whole compound}}$$

CHARLE'S LAW

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

V = volume
T = temperature

COMBINED GAS LAW

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

P = pressure
V = volume
T = temperature

DILUTION

$$M_1 V_1 = M_2 V_2$$

M = molarity
V = volume

PERCENT ERROR

$$\frac{\text{Actual value} - \text{Measured value}}{\text{Actual value}} \times 100\%$$

MOLE \rightarrow MASS

$$\text{mole} \times \text{mm} = \text{grams}$$

MASS \rightarrow MOLE

$$\frac{\text{mass}}{\text{mm}} = \text{mole}$$

TITRATION

$$N_a V_a m_a = N_b V_b m_b$$

V = volume

M = molarity

N = moles

GAY-LUSSAC'S LAW

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

P = pressure
T = temperature

Boyle's Law

$$P_1 V_1 = P_2 V_2$$

P = pressure
V = volume

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L Darling-Hammond



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