

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate K_a and K_b for weak acids and bases and pK_a and pK_b for weak acids and bases.
3. Calculate pH for weak acid (or base) solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|--------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. autoionization | 5. K_a | 6. K_b |
| | 7. pK_a | 8. pK_b |

Key Equations

- | | |
|--|--|
| 1. $K_a = \frac{[H_3O^+][A^-]}{[HA]}$ (Eq. 14.1) | 2. $K_b = \frac{[OH^-][BH]}{[B]}$ (Eq. 14.2) |
| 3. $pH = -\log[H_3O^+]$ (Eq. 14.3) | 4. $pK_a = -\log K_a$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the species indicated that are associated in the reaction. Also indicate which species acts as the conjugate acid/base.

1. H_2O and H_3O^+ are _____, where H_2O is the _____.
2. H_2O and OH^- are _____, where H_2O is the _____.
3. H_2O and H_2O are _____, where H_2O is the _____.
4. H_2O and H_2O are _____, where H_2O is the _____.
5. H_2O and H_2O are _____, where H_2O is the _____.
6. H_2O and H_2O are _____, where H_2O is the _____.
7. H_2O and H_2O are _____, where H_2O is the _____.
8. H_2O and H_2O are _____, where H_2O is the _____.
9. H_2O and H_2O are _____, where H_2O is the _____.
10. H_2O and H_2O are _____, where H_2O is the _____.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

19 2 Hydrogen Ions And Acidity Answer Key

Mark Steyvers



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Table of Contents 19 2 Hydrogen Ions And Acidity Answer Key

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Answer Key
 - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Hydrogen Ions And Acidity Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Answer Key

- Personalized Recommendations
 - 19 2 Hydrogen Ions And Acidity Answer Key User Reviews and Ratings
 - 19 2 Hydrogen Ions And Acidity Answer Key and Bestseller Lists
5. Accessing 19 2 Hydrogen Ions And Acidity Answer Key Free and Paid eBooks
 - 19 2 Hydrogen Ions And Acidity Answer Key Public Domain eBooks
 - 19 2 Hydrogen Ions And Acidity Answer Key eBook Subscription Services
 - 19 2 Hydrogen Ions And Acidity Answer Key Budget-Friendly Options
 6. Navigating 19 2 Hydrogen Ions And Acidity Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 2 Hydrogen Ions And Acidity Answer Key Compatibility with Devices
 - 19 2 Hydrogen Ions And Acidity Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Answer Key
 - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Answer Key
 - Interactive Elements 19 2 Hydrogen Ions And Acidity Answer Key
 8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Answer Key
 9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Answer Key
 - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Answer Key

- Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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