

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa from [OH⁻] and [H⁺].
3. Calculate the pH of weak (or strong) acids.

Vocabulary

- | | | |
|-------------------|-------------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium | 5. hydroxide ion (OH ⁻) | 6. autoionization |
| | 7. acid solution | 8. pH |

Key Equations

- | | |
|---|-------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H_3O^+]$ |
| 3. $pH = -\log[H_3O^+]$ | 4. $pOH = -\log[OH^-]$ |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair represented in the reaction. Also indicate which member of the pair is the stronger acid.

1. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/strong base)
2. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/weak base)
3. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/strong base)
4. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/strong base)
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9. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/strong base)
10. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (strong acid/strong base)

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10. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$

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