

# Chemistry Study Guide

Periodic Table of the Elements

1	2											3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1	H	He																							He		
2	Li	Be											B	C	N	O	F	Ne									
3	Na	Mg											Al	Si	P	S	Cl	Ar									
4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr									
5	Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe									
6	Cs	Ba			Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn								
7	Fr	Ra			Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Fl	Uup	Lv	Uus	Uuo								
			La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu										
			Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr										

## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🎯

## Matter:

Anything that has **mass** and **occupies space** 💡🌍

Examples: air, gases, liquids, solidz!

## Atoms:

Atoms are the **smallest, indivisible particle** 🔬👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\frac{1}{1836}$	-1	Orbit/Shell

# Section 2 nuclear Chemistry Study Guide Key

**John A. Olmsted, Gregory M. Williams**

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