



# Metex Multimeter User Guide

**Robert Seamans**



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web describe redox reactions in terms of electron transfer recall that oxidation involves loss of electrons and reduction involves gain of electrons represent simple redox reactions using half equations identify oxidising and reducing agents for simple redox reactions understand the reactivity of metals and metal displacement reaction

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23  $\text{Fe}^{2+} + \text{CO}_2 \rightarrow \text{Fe} + \text{CO}$  oxidation  $\text{CO} + \text{CO}_2 + 2\text{e}^-$  reduction  $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$  24  $3\text{Ag} + \text{Ni}^{3+} \rightarrow 3\text{Ag} + \text{Ni}$  oxidation reduction 25  $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$  oxidation reduction 26  $\text{O}_2$

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