

## Chemistry:

### CHAPTER 1

- Charge of electron  $\approx -1.6 \times 10^{-19}$ ; Avogadro's #  $\approx 6 \times 10^{23}$
- Mass number (A) on top of element, atomic number (Z) below
- $E = \frac{hc}{\lambda} = -R \left[ \frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$ ;  $E = qV$  (energy = charge on e- x electron volts);  $c = \lambda \nu$  ( $c = 3 \times 10^8$ )
- **Heisenberg uncertainty** – impossible to determine momentum and position simultaneously
- **Pauli exclusion** – no electrons have same quantum #'s
- Max e- in a shell  $\approx 2n^2$ , max e- in subshell  $\approx 4l+2$ 
  - Possible values for l  $\approx 0$  to  $n-1$ , possible values for ml  $\approx -l$  to  $+l$
- s can hold 2 e-, p  $\approx 6$ , d  $\approx 10$ , f  $\approx 14$
- **Aufbau principle** (building up) – lowest  $n+l$  will fill first
- For anions add extra e-, for cations remove e- from subshells with highest n first
- **Hund's** – half filled orbitals with parallel spins (bus seats)
- **Paramagnetic** – unpaired parallel spin e-, attracted to magnetic field
  - **Diamagnetic** – paired e-, repelled by magnetic field
  - **Ferromagnetic** – unpaired e-, randomly oriented dipoles, become magnetized in field
- A faraday is equal to one mole of electric charge

### CHAPTER 2

- **Z eff** increases from left to right, constant up and down
- **Ionic radius** is smallest near metalloid line and increases out
- F is smallest, most EN, highest IE and EA; Cs is the largest and least EN
- Complex ions from transition metals associate with water to form **hydration complexes**
- **Atomic radius** has the opposite trend of all others



### CHAPTER 3

- C, N, O, F, Na, Mg always abide by octet rule; H, He, Li, Be, B can have incomplete; period 3 or higher can have expanded
- **Coordinate covalent bond** – both e- are contributed by one atom (lone pair attack, acid-base)
- Ionic  $\approx$  EN difference greater than 1.7, polar covalent  $\approx$  0.5 to 1.7, nonpolar  $\approx$  less than 0.5
  - Ionic are stronger than covalent (which are both stronger than IM forces)
    - Van der Waals occurs more with close proximity and large molecules
- **Dipole moment**  $\approx p = qd$
- Triple bonds are shorter and have more energy
- Major resonance contributor is more stable (less formal charge, negative on more EN atom)
  - **Formal charge = valence – nonbonding –  $\frac{1}{2}$  bonding**
- Electron geometry is when bonding and lone are equal, molecular geo is when lone repel more

# Section 2 nuclear Chemistry Study Guide Key

**Yan Bai**



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web wind turbine power nominal power yield annual yield rotation magnitude distance repowering wind turbine power calculator calculates the power of a wind turbine from size wind speed and air density the radius is the length of a rotor blade the wind speed refers to one point in time not to an average speed

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web wind turbine blade calculator design your wind turbine blades using our software match blades to your existing generators rpm and power output customize the blade radius number and tsr to find power output for your average wind speed purchase plans for turbine blades after your design is complete