

## 17.5

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations of acids and bases and vice versa.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                              |                    |
|-------------------|------------------------------|--------------------|
| 1. autoionization | 2. conjugate acid of a base  | 3. pH              |
| 4. autoionization | 5. conjugate base of an acid | 6. polyprotic acid |
|                   | 7. weak acid                 |                    |

## Key Equations

- |                                    |                                     |
|------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $[\text{H}^+] = 10^{-\text{pH}}$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $[\text{H}^+] = 10^{-\text{pH}}$ |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid and base indicated in the reaction. Also indicate which species is the stronger acid.

1.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_2\text{O}$ )
2.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_3\text{O}^+$ )
3.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{OH}^-$ )
4.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_2\text{O}$ )
5.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{OH}^-$ )
6.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_3\text{O}^+$ )
7.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_2\text{O}$ )
8.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{OH}^-$ )
9.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_3\text{O}^+$ )
10.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_2\text{O}$ )
11.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{OH}^-$ )
12.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_3\text{O}^+$ )
13.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_2\text{O}$ )
14.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{OH}^-$ )
15.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_3\text{O}^+$ )
16.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_2\text{O}$ )
17.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{OH}^-$ )
18.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_3\text{O}^+$ )
19.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{H}_2\text{O}$ )
20.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (stronger acid is  $\text{OH}^-$ )

## Part B: Multiple Choice

1.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
2.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
3.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
4.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
5.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
6.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
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16.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
17.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
18.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
19.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
20.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

# 19 2 Hydrogen Ions And Acidity Answer Key

**Starr,Cecie Starr**



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