

# EARTH SCIENCE

## INTRODUCTION

- COMPONENTS OF STUDY**
1. **Geology:** Both physical & historical
  2. **Meteorology:** Weather & climate
  3. **Oceanography:** The oceans
  4. **Astronomy:** What is external to the atmosphere
- THE FOUR SPHERES (See The Water Cycle)**
1. **Lithosphere:** The solid earth
  2. **Hydrosphere:** The liquid earth
  3. **Atmosphere:** The gaseous earth
  4. **Biosphere:** Where life exists
- CONCERNS IN OUR PHYSICAL ENVIRONMENT**
1. **Resources:** Water, soil, metallic & non-metallic elements, sunlight, wind, etc.
  2. **Per-capita use of metallic and non-metallic elements is 11 tons per year**
  3. **Renewable vs. non-renewable resources**
  4. **Environmental and population problems**

## MINERALS

- A **mineral** is a naturally occurring, inorganic, solid material with a definite atomic structure.
  - An **ore** is a useful metallic mineral.
- A. Atom and Forces**
1. **Atom,** in ancient Greek, means "no-cut," i.e., the Ancients believed the atom to be the smallest particle of matter.
  2. Atoms are composed of neutrons, protons and electrons.
    - a. **Atomic number:** Number of protons in nucleus
    - b. **Atomic weight:** Average weight of an atom
    - c. **Isotope:** Forms of an element with identical atomic numbers, but different numbers of neutrons in the nucleus
  3. **Force:** Strong, weak, electromagnetic and gravity
- B. Properties of Minerals**
1. **Color:** Appearance or quality of light from surface
  2. **Streak:** Color of mineral in powdered form
  3. **Hardness:** Moh's Scale (See figure below)
  4. **Scratch Test:** Higher numbered materials can scratch lower numbered materials
  5. **Shape** is determined by cleavage, crystal form and fracture
  6. **Specific Gravity**
    - a. The ratio of the weight of a mineral to the weight of an equal volume of water
    - b. **Density of water = 1 gm/cm<sup>3</sup> = 1 gm/ml.**  
 i.e., lead = 7.7, aluminum = 2.7, osmium = 22

MOH'S SCALE		
Hardness	10	MINERAL
	9	Diamond
	8	Corundum
	8	Topaz
	7	Quartz
	6	Feldspar
	5	Apatite
	4	Fluorite
	3	Calcite
	2	Gypsum
	1	Talc

- C. Mineral Groups**
1. **Silicates:** Minerals with silicon and oxygen
    - a. **Silicon** is a semi-metal which forms a pyramid-shaped structure with oxygen.
    - b. **Oxygen** forms inside single tetrahedrons (silvins), chains (single) sheets (silicas), 3-D structures (silicates and quartz)
  2. **Non-Silicates**
    - a. Make up one-fourth of continental crust
    - b. **Carbonates:** Minerals with carbon and oxygen including calcite, from which we procure limestone (roads) and marble (decorative slabs)
    - c. **Oxides:** Oxygen based solids, i.e., iron sulfides (S), sulfates (SO<sub>4</sub>), halides (Cl, F)
    - d. **Halide:** Mineral form of salt
    - e. **Gypsum:** Plaster, calcium
    - f. **Native metals:** Iron, zinc, gold, silver, nickel

## ROCK TYPES



- IGNEOUS ROCKS**
- A. Formation**
1. **Extrusive igneous rocks** are formed in volcanic eruptions at surface.
  2. **Intrusive (or plutonic) igneous rocks** are produced within the earth.
  3. **Magma**
    - a. Consists mainly of silicate materials
    - b. Has gases, such as water vapor
    - c. Originates from as deep as 200 km
    - d. Differs in rate of cooling, chemical composition, and amount of gases
    - e. Produces intrusive (plutonic) igneous rocks
  4. **Lava**
    - a. Similar to magma
    - b. Most gaseous elements have escaped
    - c. Produces extrusive (volcanic) igneous rocks
- B. Structure**
1. **Crystallization (ordered pattern of ions)**
    - a. The rate of cooling strongly affects crystal size
    - b. Slow cooling = large crystals
    - c. Quick cooling = small crystals
  2. **Amorphous (unordered pattern of ions)**
    - a. Occurs with instant cooling (quenching)
    - b. Produces rocks referred to as glass
- C. Classifying Igneous Rocks**
1. **By texture, size and arrangement of interlocking crystals**
    - a. **fine-grained**
      - i. Form at surface or within upper crust
      - ii. Openings left by gas bubbles are called vesicles
    - b. **coarse-grained:** Formed deep within the interior
    - c. **porphyritic:** Large crystals imbedded in matrix of smaller crystals
  2. **By mineral composition**
    - a. Depends on composition of magma and cooling
    - b. Same magma can create rocks of varying mineral content.
    - c. As magma cools, certain minerals crystallize first (at very high temperatures).

- METAMORPHIC ROCKS**
- A. Formation**
1. **Metamorphic:** To change form
  2. **Regional metamorphism:** Material under intense stresses, high temperatures
  3. **Contact (thermal) metamorphism:** Changes caused by proximity to magma
  4. **Low grade metamorphism:** Slate becomes shale
  5. **High grade metamorphism:** Fossils become rock
  6. **Cliffs form** during mountain building
  7. Most metamorphic rocks are harder than sedimentary rocks
  8. **Metamorphism changes texture**
  9. **Low-grade metamorphism** makes material more compact, more dense
  10. **Foliated texture:** Particles of material are brought into line with one another
- B. Agents of Metamorphism**
1. **Heat**
    - a. Most important agent
    - b. Provides energy for chemical reactions
    - c. For example: Clay recrystallizes into a mineral at great temperature
  2. **Pressure**
  3. **Chemical Activity:** Most common chemical agent in water
- C. Types of Metamorphic Rocks**
1. **Gneiss:** Fine grained, dark flinty rock with randomly arranged minerals
  2. **Slate:** Fine grained, often gray, foliated rock split easily along cleavage, planes of mica flakes aligned by pressure
  3. **Marble:** Granular or sugary-textured rock formed from limestone
  4. **Phyllite:** Silky, foliated rock more coarsely grained than slate
  5. **Schist:** foliated rock, more coarsely grained and of higher metamorphic grade than phyllite
  6. **Gneiss:** Foliated, banded rock of the highest metamorphic grade
  7. **Quartzite:** Very hard, granular quartz rock, formed from sandstone

- BOWEN'S REACTION SERIES**
- If a mineral, which has already formed, remains in the magma, it will react with the remaining magma to produce the next mineral in the sequence. For example, *olivine forms first; olivine then reacts with remaining magma to form pyroxene.*
- A. Right Side of the Bowen Series**
1. Calcium rich parts of the magma form small crystals of feldspar.
  2. These react with sodium in the magma to become more and more sodium rich.
- B. End of Cooling**
1. When everything is almost cool, remaining magma will have high silicose content, and quartz will form.
  2. When cooling is complete, everything that cooled at the same time will usually be close to one another (feldspar, mica and quartz cool near one another to make granite).



# Earth Science High School Essential Question

**VM Jensen**



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Teaching Discipline-Specific Literacies in Grades 6-12 Vicky I. Zygouris-Coe, 2014-10-30 Comprehensive timely and relevant this text offers an approach to discipline specific literacy instruction that is aligned with the Common Core State Standards and the needs of teachers students and secondary schools across the nation It is essential that teachers know how to provide instruction that both develops content and literacy knowledge and skills and aims at reducing student achievement gaps Building on the research supported premise that discipline specific reading instruction is key to achieving these goals this text provides practical guidance and strategies for prospective and practicing content area teachers and other educators on how to prepare all students to succeed in college and the workforce Pedagogical features in each chapter engage readers in digging deeper and in applying the ideas and strategies presented in their own contexts Classroom Life real 6 12 classroom scenarios and interviews with content area teachers Common Core State Standards Connections College Career and Workforce Connections Applying Discipline Specific Literacies Think Like an Expert habits of thinking and learning specific to each discipline Digital Literacies Differentiating Instruction Reflect and Apply Questions Extending Learning Activities The Companion Website includes Lesson plan resources Annotated links to video files Annotated links to additional resources and information Glossary Flashcards For Instructors All images and figures used in the text provided in an easily downloadable format For Instructors PowerPoint lecture slides

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**Proceedings of the High School Conference of ...** Horace Adelbert Hollister, 1919

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