

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and an appropriate equilibrium.
2. Calculate hydroxide ion concentration, hydronium ion concentration, and pH of an acid or base solution.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                             |                  |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion  | 5. hydronium ion            | 6. base solution |
|                   | 7. weak acid                | 8. pH            |

## Key Equations

- |  |  |
|--|--|
| 1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ | 2. $\text{pH} = -\log[\text{H}_3\text{O}^+]$                         |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$  | 4. $\text{pH} = \text{pK}_a + \log \frac{[\text{A}^-]}{[\text{HA}]}$ |

## Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions. Show calculations for all problems.

1. What products are formed when hydrochloric acid and methanol react? Write the balanced chemical equation for this reaction and identify the products.
2. Hydrochloric acid ionizes in water. Write the chemical equation for this reaction and identify the products.
3. Write the chemical equation for the autoionization of water.
4. Write the chemical equation for the autoionization of water and identify the products.
5. Write the chemical equation for the autoionization of water and identify the products.
6. Write the chemical equation for the autoionization of water and identify the products.
7. Write the chemical equation for the autoionization of water and identify the products.

8.  $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$
9.  $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$
10.  $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$
11.  $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$
12.  $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$
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14.  $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$
15.  $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$
16.  $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$

# 19 2 Section Review Hydrogen Ions And Acidity

**RJ Shavelson**



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