

# Acids & Bases HOMEWORK

## Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- HCl **hydrochloric acid**
- H<sub>2</sub>SO<sub>4</sub> **sulfuric acid**
- HNO<sub>3</sub> **nitric acid**
- HClO<sub>4</sub> **perchloric acid**
- Strontium hydroxide **LiOH**
- potassium hydroxide **KOH**
- calcium hydroxide **Ca(OH)<sub>2</sub>**
- barium hydroxide **Ca(OH)<sub>2</sub>**
- ammonium hydroxide **Li(OH)<sub>4</sub>**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.  
HClO<sub>4</sub> is STRONG with a lower cell completely dissociate in water

- HCl**      **H<sub>2</sub>**      **H<sub>2</sub>O**      **Mg(OH)<sub>2</sub>**      **HNO<sub>3</sub>**      **Ca(OH)<sub>2</sub>**      **HNO<sub>2</sub>**      **H<sub>2</sub>SO<sub>4</sub>**

3. What is the formula for hydrosulfuric acid? **H<sub>2</sub>S**

4.  $\text{BaCO}_3$  will react with water and lose a  $\text{H}^+$  ion in the process.

- Write a balanced equation for the reaction.  **$\text{BaCO}_3 + \text{H}_2\text{O} \rightleftharpoons \text{BaCO}_3^+ + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{BaCO}_3$  is a weak acid.**
- Would you define  $\text{BaCO}_3$  as a proton donor or a proton acceptor? **proton donor, because it is losing a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{BaCO}_3$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
- Identify the conjugate base of  $\text{BaCO}_3$ .  **$\text{BaCO}_3^+$**
- Identify  $\text{BaCO}_3$  as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's/atoms.**

5.  $\text{F}^-$  will react with water and gain a  $\text{H}^+$  ion in the process.

- Write a balanced equation for the reaction.  **$\text{F}^- + \text{H}_2\text{O} \rightleftharpoons \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{F}^-$  is a weak base.**
- Would you define  $\text{F}^-$  as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{F}^-$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of  $\text{F}^-$ . **HF**
- Identify  $\text{F}^-$  as polyprotic, amphoteric, or neither. **neither**

6.  $\text{HCO}_3^-$  will react with water and lose a  $\text{H}^+$  ion in the process.

- Write a balanced equation for the reaction.  **$\text{HCO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{HCO}_3^{2-} + \text{H}_3\text{O}^+$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because  $\text{HCO}_3^-$  is a weak acid.**
- Would you define  $\text{HCO}_3^-$  as a proton donor or a proton acceptor? **proton donor, because it is losing a  $\text{H}^+$  ion.**
- In your written equation, is  $\text{HCO}_3^-$  acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
- Identify the conjugate base of  $\text{HCO}_3^-$ .  **$\text{CO}_3^{2-}$**
- Identify  $\text{HCO}_3^-$  as polyprotic, amphoteric, or neither. **amphoteric, because of both a negative charge (only will take a proton) and a H's/atoms (can act like an acid).**

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