

Chemistry 2nd Semester Final Review **KEY**

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

1. If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)
 $V_1=3.8 \text{ L}; P_1=2.71 \text{ atm}$
 $V_2=1.47 \text{ L}; P_2=? \text{ Atm}$
Since all units are standard (L, atm), we do not need to convert the units before proceeding.
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
T goes away (not mentioned in problem)
 $P_1V_1=P_2V_2$
Isolate for P₂: $P_2=(P_1V_1)/V_2$
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})=7.0 \text{ atm}$ (L cancelled and it's 2 sig figs)
2. A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)
 $V_1=1.49 \text{ L}, T_1=34.75 \text{ degrees Celsius}$
 $V_2=? \text{ L}, T_2=78.41 \text{ degrees Celsius}$
We need to convert from Celsius to Kelvins before proceeding.
 $T_1=34.75 + 273=307.75 \text{ K}$
 $T_2=78.41 + 273=351.41 \text{ K}$
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
Get rid of P
 $V_1/T_1=V_2/T_2$
Isolate V₂: $V_2=(V_1/T_1)T_2$
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})=1.70 \text{ L}$
3. What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?
The combined gas law won't work here. You only have one condition and nothing changes.
 $PV=nRT$ will work. (n= moles and R= 0.08206 L atm/K mol)
 $P=1.04 \text{ atm}, V=? \text{ L}, n=?$ OOPS we need to convert from grams to moles!, $T=84.7 \text{ degrees Celsius} + 273=357.7 \text{ K}$
n: Hydrogen gas is diatomic!
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/2.02 \text{ g H}_2)=4.193069307 \text{ mol H}_2$
Isolate V: $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K})/1.04 \text{ atm}=118 \text{ L}$
4. What volume is occupied by 56.75 g of oxygen gas at STP?

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Annelies Wilder-Smith



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Table of Contents 2nd Semester Chemistry Final Exam Study Guide

1. Understanding the eBook 2nd Semester Chemistry Final Exam Study Guide
 - The Rise of Digital Reading 2nd Semester Chemistry Final Exam Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying 2nd Semester Chemistry Final Exam Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2nd Semester Chemistry Final Exam Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2nd Semester Chemistry Final Exam Study Guide
 - Personalized Recommendations
 - 2nd Semester Chemistry Final Exam Study Guide User Reviews and Ratings

- 2nd Semester Chemistry Final Exam Study Guide and Bestseller Lists
- 5. Accessing 2nd Semester Chemistry Final Exam Study Guide Free and Paid eBooks
 - 2nd Semester Chemistry Final Exam Study Guide Public Domain eBooks
 - 2nd Semester Chemistry Final Exam Study Guide eBook Subscription Services
 - 2nd Semester Chemistry Final Exam Study Guide Budget-Friendly Options
- 6. Navigating 2nd Semester Chemistry Final Exam Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - 2nd Semester Chemistry Final Exam Study Guide Compatibility with Devices
 - 2nd Semester Chemistry Final Exam Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2nd Semester Chemistry Final Exam Study Guide
 - Highlighting and Note-Taking 2nd Semester Chemistry Final Exam Study Guide
 - Interactive Elements 2nd Semester Chemistry Final Exam Study Guide
- 8. Staying Engaged with 2nd Semester Chemistry Final Exam Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2nd Semester Chemistry Final Exam Study Guide
- 9. Balancing eBooks and Physical Books 2nd Semester Chemistry Final Exam Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2nd Semester Chemistry Final Exam Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 2nd Semester Chemistry Final Exam Study Guide
 - Setting Reading Goals 2nd Semester Chemistry Final Exam Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 2nd Semester Chemistry Final Exam Study Guide
 - Fact-Checking eBook Content of 2nd Semester Chemistry Final Exam Study Guide
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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