

## Chemistry:

### CHAPTER 1

- Charge of electron  $\approx -1.6 \times 10^{-19}$ ; Avogadro's #  $\approx 6 \times 10^{23}$
- Mass number (A) on top of element, atomic number (Z) below
- $E = \frac{hc}{\lambda} = -R \left[ \frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$ ;  $E = qV$  (energy = charge on e- x electron volts);  $c = \lambda \nu$  ( $c = 3 \times 10^8$ )
- **Heisenberg uncertainty** – impossible to determine momentum and position simultaneously
- **Pauli exclusion** – no electrons have same quantum #'s
- Max e- in a shell  $\approx 2n^2$ , max e- in subshell  $\approx 4l+2$ 
  - Possible values for l  $\approx 0$  to  $n-1$ , possible values for ml  $\approx -l$  to  $+l$
- s can hold 2 e-, p  $\approx 6$ , d  $\approx 10$ , f  $\approx 14$
- **Aufbau principle** (building up) – lowest  $n+l$  will fill first
- For anions add extra e-, for cations remove e- from subshells with highest n first
- **Hund's** – half filled orbitals with parallel spins (bus seats)
- **Paramagnetic** – unpaired parallel spin e-, attracted to magnetic field
  - **Diamagnetic** – paired e-, repelled by magnetic field
  - **Ferromagnetic** – unpaired e-, randomly oriented dipoles, become magnetized in field
- A faraday is equal to one mole of electric charge

### CHAPTER 2

- **Z eff** increases from left to right, constant up and down
- **Ionic radius** is smallest near metalloid line and increases out
- F is smallest, most EN, highest IE and EA; Cs is the largest and least EN
- Complex ions from transition metals associate with water to form **hydration complexes**
- **Atomic radius** has the opposite trend of all others



### CHAPTER 3

- C, N, O, F, Na, Mg always abide by octet rule; H, He, Li, Be, B can have incomplete; period 3 or higher can have expanded
- **Coordinate covalent bond** – both e- are contributed by one atom (lone pair attack, acid-base)
- Ionic  $\approx$  EN difference greater than 1.7, polar covalent  $\approx$  0.5 to 1.7, nonpolar  $\approx$  less than 0.5
  - Ionic are stronger than covalent (which are both stronger than IM forces)
    - Van der Waals occurs more with close proximity and large molecules
- **Dipole moment**  $\approx p = qd$
- Triple bonds are shorter and have more energy
- Major resonance contributor is more stable (less formal charge, negative on more EN atom)
  - **Formal charge = valence – nonbonding – ½ bonding**
- Electron geometry is when bonding and lone are equal, molecular geo is when lone repel more

# Section 2 nuclear Chemistry Study Guide Key

**WJ Hussar**



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## **Table of Contents Section 2nuclear Chemistry Study Guide Key**

1. Understanding the eBook Section 2nuclear Chemistry Study Guide Key
  - The Rise of Digital Reading Section 2nuclear Chemistry Study Guide Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Section 2nuclear Chemistry Study Guide Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section 2nuclear Chemistry Study Guide Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Section 2nuclear Chemistry Study Guide Key
  - Personalized Recommendations
  - Section 2nuclear Chemistry Study Guide Key User Reviews and Ratings
  - Section 2nuclear Chemistry Study Guide Key and Bestseller Lists
5. Accessing Section 2nuclear Chemistry Study Guide Key Free and Paid eBooks

- Section 2 nuclear Chemistry Study Guide Key Public Domain eBooks
  - Section 2 nuclear Chemistry Study Guide Key eBook Subscription Services
  - Section 2 nuclear Chemistry Study Guide Key Budget-Friendly Options
6. Navigating Section 2 nuclear Chemistry Study Guide Key eBook Formats
    - ePub, PDF, MOBI, and More
    - Section 2 nuclear Chemistry Study Guide Key Compatibility with Devices
    - Section 2 nuclear Chemistry Study Guide Key Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Section 2 nuclear Chemistry Study Guide Key
    - Highlighting and Note-Taking Section 2 nuclear Chemistry Study Guide Key
    - Interactive Elements Section 2 nuclear Chemistry Study Guide Key
  8. Staying Engaged with Section 2 nuclear Chemistry Study Guide Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Section 2 nuclear Chemistry Study Guide Key
  9. Balancing eBooks and Physical Books Section 2 nuclear Chemistry Study Guide Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Section 2 nuclear Chemistry Study Guide Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Section 2 nuclear Chemistry Study Guide Key
    - Setting Reading Goals Section 2 nuclear Chemistry Study Guide Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Section 2 nuclear Chemistry Study Guide Key
    - Fact-Checking eBook Content of Section 2 nuclear Chemistry Study Guide Key
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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web the night gardener interest age 4 9 reading age 5 4 5 years 6 8 years 9 11 years adoption environment friendship nature the night gardener author the fan brothers publisher frances lincoln william looks out of his window in the orphanage on grimloch lane one day to find a commotion going on outside

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web mar 16 2020 written in a clearly gothic style the action largely revolves around the decidedly spooky night gardener a

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