

Unit 8 Review

1. The balanced equation  $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$  tells us that 4.0 mol  $\text{H}_2$   
 A) reacts with 2.0 mol  $\text{P}_4$   
 B) produces 8.0 mol  $\text{PH}_3$   
 C) cannot react with phosphorus  
 D) produces 2.7 mol  $\text{PH}_3$   
 E) reacts with 4.0 mol  $\text{P}_4$

2. A 1.5-mol sample of  $\text{Fe}_2\text{O}_3$  was decomposed according to the equation:  
 $2\text{Fe}_2\text{O}_3(\text{s}) \rightarrow 2\text{Fe}_2\text{O}(\text{s}) + 3\text{O}_2(\text{g})$

How many moles of  $\text{O}_2$  are formed assuming 100% yield?

- A) 0.9 mol
- B) 1.5 mol
- C) 1.2 mol
- D) 0.7 mol
- E) 2.0 mol

3. The rusting of iron is represented by the equation  $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$ . If you have a 1.00-mol sample of iron, how many moles of  $\text{Fe}_2\text{O}_3$  will there be after the iron has rusted completely?

- A) 0.527 mol
- B) 0.750 mol
- C) 1.00 mol
- D) 1.50 mol
- E) 2.07 mol

4. Consider the following reaction, where X represents an unknown element.  
 $8\text{X}_2(\text{s}) + 2\text{B}_2\text{O}_3(\text{s}) \rightarrow \text{B}_2\text{X}_4(\text{s}) + 3\text{X}_2\text{O}_3(\text{g})$

If 175 g of X reacts completely with diboron trioxide to produce 1.43 mol of  $\text{B}_2\text{X}_4$ , what is the identity of X?

- A) Ge
- B) Mg
- C) Si
- D) Sn
- E) C

5. Refer to the following equation:  
 $4\text{NH}_3(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{NO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$

How many molecules of  $\text{NO}_2$  are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08
- B)  $3.84 \times 10^{23}$
- C)  $1.82 \times 10^{23}$
- D) 1.39
- E) none of these

6. In the reaction  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ , how many moles of ammonia would be produced from 1.00 mol of hydrogen and excess nitrogen?

- A) 1.42 mol
- B) 3.37 mol
- C) 0.727 mol
- D) 2.18 mol
- E) 0.343 mol

7. Refer to the following unbalanced equation:  
 $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

What mass of oxygen ( $\text{O}_2$ ) is required to react completely with 13.1 g of  $\text{C}_2\text{H}_6$ ?

- A)  $2.98 \times 10^2$  g
- B) 23.1 g
- C) 4.86 g
- D) 49.2 g
- E) 0.192 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A)  $1.37 \times 10^{23}$  molecules
- B)  $1.00 \times 10^{23}$  molecules
- C)  $4.14 \times 10^{23}$  molecules
- D)  $3.10 \times 10^{23}$  molecules
- E)  $5.17 \times 10^{23}$  molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane,  $\text{C}_8\text{H}_{18}$ , in the following reaction.



- A) 79.2 g  $\text{CO}_2$
- B) 158 g  $\text{CO}_2$
- C) 89.1 g  $\text{CO}_2$
- D) 1.258 g  $\text{CO}_2$
- E) 1.400 g  $\text{CO}_2$

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction.



- A)  $1.03 \times 10^{23}$  molecules  $\text{O}_2$
- B)  $2.28 \times 10^{23}$  molecules  $\text{O}_2$
- C)  $1.71 \times 10^{23}$  molecules  $\text{O}_2$
- D)  $3.42 \times 10^{23}$  molecules  $\text{O}_2$
- E)  $5.13 \times 10^{23}$  molecules  $\text{O}_2$

# Multiple Choice Stoichiometry Test

**J Spring**



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**Roadmap to the Regents** Sasha Alcott, 2003 If Students Need to Know It It's in This Book This book develops the chemistry skills of high school students It builds skills that will help them succeed in school and on the New York Regents Exams Why The Princeton Review We have more than twenty years of experience helping students master the skills needed to excel on standardized tests Each year we help more than 2 million students score higher and earn better grades We Know the New York Regents Exams Our experts at The Princeton Review have analyzed the New York Regents Exams and this book provides the most up to date thoroughly researched practice possible We break down the test into individual skills to familiarize students with the test's structure while increasing their overall skill level We Get Results We know what it takes to succeed in the classroom and on tests This book includes strategies that are proven to improve student performance We provide a breakdown of the skills based on New York standards and objectives hundreds of practice questions organized by skill two complete practice New York Regents Exams in Physical Setting Chemistry

**CLEP Success** Peterson's, 2007-05-04 Offers advice on improving scores on the CLEP and includes reviews for five CLEP subjects and five full length practice tests with explanatory answers

**Master the CLEP** Peterson's, 2012-04-30 Peterson's Master the CLEP is the essential prep guide for non traditional students who are looking to begin a college career and for current students who want to save on tuition costs by testing out of certain courses When you use this guide you will get a basic overview of the CLEP a look at the general examinations and the vast number of subject specific exams tips for registering and preparing for the CLEP and a guide to interpreting your test scores This is followed by extensive reviews of the five CLEP general examinations with pre and post tests and subject matter overviews that will help you pinpoint your strengths and weaknesses The Appendix will go into detail about a number of CLEP subject exams

**Journal of Chemical Education**, 1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers

*The ETS Test Collection Catalog* Educational Testing Service. Test Collection, 1993 The major source of information on the availability of standardized tests Wilson Library Bulletin Covers commercially available standardized tests and hard to locate research instruments

**Optimizing STEM Education With Advanced ICTs and Simulations** Levin, Ilya, Tsybulsky, Dina, 2017-06-05 The role of technology in educational settings has become increasingly prominent in recent years When utilized effectively these tools provide a higher quality of learning for students Optimizing STEM Education With Advanced ICTs and Simulations is an innovative reference source for the latest scholarly research on the integration of digital tools for enhanced STEM based learning environments

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*Tests* Richard C. Sweetland, Daniel J. Keyser, 1983 Descriptions of over 3000 tests in English intended as a guide for psychologists educators and other personnel who need test information to meet their assessment needs Topical arrangement under 3 main sections of psychology education and business Each entry gives test name associated personal names intended age group purpose description time range scoring cost and publisher Indexes by test titles authors publishers visually impaired tests and scoring services

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## **Table of Contents Multiple Choice Stoichiometry Test**

1. Understanding the eBook Multiple Choice Stoichiometry Test
  - The Rise of Digital Reading Multiple Choice Stoichiometry Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Stoichiometry Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Stoichiometry Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Stoichiometry Test
  - Personalized Recommendations
  - Multiple Choice Stoichiometry Test User Reviews and Ratings
  - Multiple Choice Stoichiometry Test and Bestseller Lists
5. Accessing Multiple Choice Stoichiometry Test Free and Paid eBooks
  - Multiple Choice Stoichiometry Test Public Domain eBooks
  - Multiple Choice Stoichiometry Test eBook Subscription Services
  - Multiple Choice Stoichiometry Test Budget-Friendly Options

6. Navigating Multiple Choice Stoichiometry Test eBook Formats
  - ePub, PDF, MOBI, and More
  - Multiple Choice Stoichiometry Test Compatibility with Devices
  - Multiple Choice Stoichiometry Test Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Multiple Choice Stoichiometry Test
  - Highlighting and Note-Taking Multiple Choice Stoichiometry Test
  - Interactive Elements Multiple Choice Stoichiometry Test
8. Staying Engaged with Multiple Choice Stoichiometry Test
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Multiple Choice Stoichiometry Test
9. Balancing eBooks and Physical Books Multiple Choice Stoichiometry Test
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Multiple Choice Stoichiometry Test
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Stoichiometry Test
  - Setting Reading Goals Multiple Choice Stoichiometry Test
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Stoichiometry Test
  - Fact-Checking eBook Content of Multiple Choice Stoichiometry Test
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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