

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations from K_w and K_a or K_b and the concentration of a weak acid or base.
3. Calculate K_a or K_b from pH or pOH.

Vocabulary

- | | | |
|-------------------|-----------------------------|-----------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid ionization constant |
| 4. hydroxide ion | 5. hydronium ion | 6. base ionization constant |
| | 7. weak acid | 8. pH |

Key Equations

- | | |
|--|---|
| 1. $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ | 2. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ | 4. $\text{pOH} = -\log[\text{OH}^-]$ |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second column with the number in the first column.

1. H_2O and OH^- are a conjugate acid-base pair.
2. H_2O and H_3O^+ are a conjugate acid-base pair.
3. H_2O and H^+ are a conjugate acid-base pair.
4. H_2O and H^+ are a conjugate acid-base pair.
5. H_2O and H^+ are a conjugate acid-base pair.
6. H_2O and H^+ are a conjugate acid-base pair.
7. H_2O and H^+ are a conjugate acid-base pair.
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19. H_2O and H^+ are a conjugate acid-base pair.
20. H_2O and H^+ are a conjugate acid-base pair.

1. H_2O and OH^-

2. H_2O and H_3O^+

3. H_2O and H^+

4. H_2O and H^+

5. H_2O and H^+

6. H_2O and H^+

7. H_2O and H^+

8. H_2O and H^+

9. H_2O and H^+

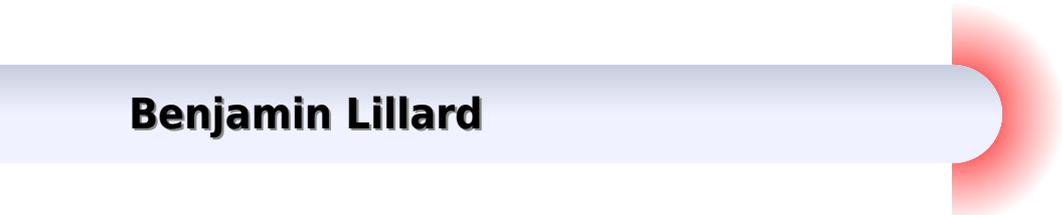
10. H_2O and H^+

11. H_2O and H^+

12. H_2O and H^+

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