

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode $T \gg \theta_M$) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute $\frac{3}{2}R$ to the molar heat capacity. If we denote the number of active rotational modes by ν_R^\ddagger (so for most molecules at normal temperatures $\nu_R^\ddagger = 2$ for linear molecules, and 3 for non-linear molecules), then

$$C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \dots$$

16.10 The equipartition theorem states that the average energy of a degree of freedom is $\frac{1}{2}kT$ per molecule. For a diatomic molecule, the degrees of freedom are the translational, rotational, and vibrational modes. The translational modes are always active and contribute $\frac{3}{2}kT$ to the energy. The rotational modes are active at room temperature and contribute kT to the energy. The vibrational modes are not active at room temperature and contribute $\frac{1}{2}kT$ to the energy. The total energy is $\frac{3}{2}kT + kT + \frac{1}{2}kT = \frac{5}{2}kT$. The heat capacity is $C_V = \frac{dU}{dT} = \frac{5}{2}R$.

Solutions to exercises

16.11 The vibrational energy is given by $\epsilon = \frac{1}{2}h\nu$. The vibrational temperature is $\theta_v = \frac{h\nu}{k}$. The vibrational energy is $\epsilon = \frac{1}{2}k\theta_v$.

$$C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \dots$$

At room temperature, $T \approx 300\text{ K}$, $\theta_v \approx 2000\text{ K}$, $\frac{T}{\theta_v} \approx 0.15$.

$$C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \frac{1}{2}R \left(\frac{T}{\theta_v} \right)^2 = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \frac{1}{2}R (0.15)^2$$

16.12 The vibrational energy is

$$U = \frac{3}{2}RT + \frac{\nu_R^\ddagger}{2}RT + \frac{1}{2}R \sum_{i=1}^{\nu_v} \frac{\theta_{v,i}}{T} \left(e^{-\theta_{v,i}/T} + 2e^{-2\theta_{v,i}/T} + \dots \right)$$

16.13 $C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \dots$

At room temperature, $T \approx 300\text{ K}$,

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Physico-chemical Aspects of Textile Coloration Stephen M. Burkinshaw, 2016-02-08

The production of textile materials comprises a very large and complex global industry that utilises a diverse range of fibre types and creates a variety of textile products As the great majority of such products are coloured predominantly using aqueous dyeing processes the coloration of textiles is a large scale global business in which complex procedures are used to apply different types of dye to the various types of textile material The development of such dyeing processes is the result of substantial research activity undertaken over many decades into the physico chemical aspects of dye adsorption and the establishment of dyeing theory which seeks to describe the mechanism by which dyes interact with textile fibres Physico Chemical Aspects of Textile Coloration provides a comprehensive treatment of the physical chemistry involved in the dyeing of the major types of natural man made and synthetic fibres with the principal types of dye The book covers fundamental aspects of the physical and chemical structure of both fibres and dyes together with the structure and properties of water in relation to dyeing dyeing as an area of study as well as the terminology employed in dyeing technology and science

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sql fragment to append to the main join clause t type string where sql fragment to append to the main where clause t t
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 code to run by adding this php code parse uri explode wp content server script filename require once parse uri 0 wp load php
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