

Unit 8 Review

1. The balanced equation  $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$  tells us that 4.0 mol  $\text{H}_2$

A) reacts with 2.0 mol  $\text{P}_4$ .  
 B) produces 8.0 mol  $\text{PH}_3$ .  
 C) cannot react with phosphorus.  
 D) produces 2.7 mol  $\text{PH}_3$ .  
 E) reacts with 4.0 mol  $\text{P}_4$ .

2. A 1.5-mol sample of  $\text{Fe}_2\text{O}_3$  was decomposed according to the equation:

$$2\text{Fe}_2\text{O}_3(\text{s}) \rightarrow 2\text{Fe}_2\text{O}(\text{s}) + 3\text{O}_2(\text{g})$$

How many moles of  $\text{O}_2$  are formed assuming 100% yield?

- A) 0.9 mol  
 B) 1.5 mol  
 C) 1.2 mol  
 D) 0.7 mol  
 E) 2.0 mol

3. The rusting of iron is represented by the equation  $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$ . If you have a 1.00-mol sample of iron, how many moles of  $\text{Fe}_2\text{O}_3$  will there be after the iron has rusted completely?

- A) 0.527 mol  
 B) 0.750 mol  
 C) 1.00 mol  
 D) 1.50 mol  
 E) 2.27 mol

4. Consider the following reaction, where X represents an unknown element:

$$8\text{X}_2(\text{s}) + 2\text{B}_2\text{O}_3(\text{s}) \rightarrow \text{B}_2\text{X}_6(\text{s}) + 3\text{X}_2\text{O}_3(\text{g})$$

If 175 g of X reacts completely with diboron trioxide to produce 1.43 mol of  $\text{B}_2\text{X}_6$ , what is the identity of X?

- A) Ge  
 B) Mg  
 C) Si  
 D) Sn  
 E) C

5. Refer to the following equation:

$$4\text{NH}_3(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$$

How many molecules of  $\text{NO}$  are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08  
 B)  $3.84 \times 10^{23}$   
 C)  $1.82 \times 10^{24}$   
 D) 1.39  
 E) none of these

6. In the reaction  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ , how many moles of ammonia would be produced from 1.00 mol of hydrogen and excess nitrogen?

- A) 1.42 mol  
 B) 3.27 mol  
 C) 0.727 mol  
 D) 2.18 mol  
 E) 0.343 mol

7. Refer to the following unbalanced equation:

$$\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$

What mass of oxygen ( $\text{O}_2$ ) is required to react completely with 13.1 g of  $\text{C}_2\text{H}_6$ ?

- A)  $2.98 \times 10^2$  g  
 B) 23.1 g  
 C) 4.86 g  
 D) 49.2 g  
 E) 0.192 g

8. How many molecules of carbon dioxide will be formed if 7.27 g of propane is burned in the following reaction?



- A)  $1.37 \times 10^{23}$  molecules  
 B)  $1.00 \times 10^{23}$  molecules  
 C)  $4.14 \times 10^{23}$  molecules  
 D)  $3.10 \times 10^{23}$  molecules  
 E)  $5.17 \times 10^{23}$  molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane,  $\text{C}_8\text{H}_{18}$ , in the following reaction:



- A) 79.2 g  $\text{CO}_2$   
 B) 158 g  $\text{CO}_2$   
 C) 89.1 g  $\text{CO}_2$   
 D) 1.258 g  $\text{CO}_2$   
 E) 1.400 g  $\text{CO}_2$

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction:



- A)  $1.03 \times 10^{23}$  molecules  $\text{O}_2$   
 B)  $2.28 \times 10^{23}$  molecules  $\text{O}_2$   
 C)  $1.71 \times 10^{23}$  molecules  $\text{O}_2$   
 D)  $3.42 \times 10^{23}$  molecules  $\text{O}_2$   
 E)  $5.13 \times 10^{23}$  molecules  $\text{O}_2$

# Multiple Choice Stoichiometry Test

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**Multiple Choice Stoichiometry Test:**

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### **Table of Contents Multiple Choice Stoichiometry Test**

1. Understanding the eBook Multiple Choice Stoichiometry Test
  - The Rise of Digital Reading Multiple Choice Stoichiometry Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Stoichiometry Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Stoichiometry Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Stoichiometry Test
  - Personalized Recommendations
  - Multiple Choice Stoichiometry Test User Reviews and Ratings
  - Multiple Choice Stoichiometry Test and Bestseller Lists

5. Accessing Multiple Choice Stoichiometry Test Free and Paid eBooks
  - Multiple Choice Stoichiometry Test Public Domain eBooks
  - Multiple Choice Stoichiometry Test eBook Subscription Services
  - Multiple Choice Stoichiometry Test Budget-Friendly Options
6. Navigating Multiple Choice Stoichiometry Test eBook Formats
  - ePub, PDF, MOBI, and More
  - Multiple Choice Stoichiometry Test Compatibility with Devices
  - Multiple Choice Stoichiometry Test Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Multiple Choice Stoichiometry Test
  - Highlighting and Note-Taking Multiple Choice Stoichiometry Test
  - Interactive Elements Multiple Choice Stoichiometry Test
8. Staying Engaged with Multiple Choice Stoichiometry Test
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Multiple Choice Stoichiometry Test
9. Balancing eBooks and Physical Books Multiple Choice Stoichiometry Test
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Multiple Choice Stoichiometry Test
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Stoichiometry Test
  - Setting Reading Goals Multiple Choice Stoichiometry Test
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Stoichiometry Test
  - Fact-Checking eBook Content of Multiple Choice Stoichiometry Test
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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