

## 17.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                               |                   |
|-------------------|-------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair   | 3. acid solution  |
| 4. autoionization | 5. hydronium ion ( $H_3O^+$ ) | 6. autoionization |
|                   | 7. acid solution              | 8. pH             |

## Key Equations

- |  |                                    |
|--|------------------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (Eq. 17.1) | 2. $pH = -\log[H_3O^+]$ (Eq. 17.2) |
| 3. $pH = -\log[H_3O^+]$                              | 4. $pH = -\log[H_3O^+]$            |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Write the number in the space provided.

1.  $H_2O$  and  $H_3O^+$  are \_\_\_\_\_, whereas  $H_2O$  and  $OH^-$  are \_\_\_\_\_.
2.  $H_2O$  and  $OH^-$  are \_\_\_\_\_, whereas  $H_2O$  and  $H_3O^+$  are \_\_\_\_\_.
3.  $H_2O$  and  $H_3O^+$  are \_\_\_\_\_, whereas  $H_2O$  and  $OH^-$  are \_\_\_\_\_.
4.  $H_2O$  and  $OH^-$  are \_\_\_\_\_, whereas  $H_2O$  and  $H_3O^+$  are \_\_\_\_\_.
5.  $H_2O$  and  $H_3O^+$  are \_\_\_\_\_, whereas  $H_2O$  and  $OH^-$  are \_\_\_\_\_.
6.  $H_2O$  and  $OH^-$  are \_\_\_\_\_, whereas  $H_2O$  and  $H_3O^+$  are \_\_\_\_\_.
7.  $H_2O$  and  $H_3O^+$  are \_\_\_\_\_, whereas  $H_2O$  and  $OH^-$  are \_\_\_\_\_.
8.  $H_2O$  and  $OH^-$  are \_\_\_\_\_, whereas  $H_2O$  and  $H_3O^+$  are \_\_\_\_\_.

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# 19 2 Hydrogen Ions And Acidity Answer Key

**Y Pai**



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