

# Acids & Bases HOMEWORK

## Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- HCl ~~hydrochloric acid~~
- H<sub>2</sub>SO<sub>4</sub> ~~hydrochloric acid~~
- HNO<sub>3</sub> ~~nitric acid~~
- HClO<sub>4</sub> ~~perchloric acid~~
- Strontium hydroxide ~~LiOH~~
- potassium hydroxide ~~KOH~~
- calcium hydroxide ~~CaOH~~
- calcium hydroxide ~~Ca(OH)<sub>2</sub>~~
- ammonium hydroxide ~~Li(OH)<sub>4</sub>~~

*HClO<sub>4</sub> is STRONG with a lower cell completely dissociate in water*

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

HF

H<sub>2</sub>S

Mg(OH)<sub>2</sub>

HNO<sub>3</sub>

Ca(OH)<sub>2</sub>

H<sub>2</sub>SO<sub>4</sub>

HClO<sub>4</sub>

3. What is the formula for hydrosulfuric acid? H<sub>2</sub>S

4. H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> will react with water and lose a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction.  $H_2C_2O_4 + H_2O \rightleftharpoons HC_2O_4^- + H_3O^+$
- Would you expect a large or small equilibrium constant for this reaction? small, because H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is a weak acid.
- Would you define H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> as a proton donor or a proton acceptor? proton donor, because it is losing a H<sup>+</sup> ion.
- In your written equation, is H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? Brønsted-Lowry acid, because it is losing a proton.
- Identify the conjugate base of H<sub>2</sub>C<sub>2</sub>O<sub>4</sub>. HC<sub>2</sub>O<sub>4</sub><sup>-</sup>
- Identify H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> as polyprotic, amphoteric, or neither. polyprotic, because it contains multiple H's to give.

5. F<sup>-</sup> will react with water and gain a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction.  $F^- + H_2O \rightleftharpoons HF + OH^-$
- Would you expect a large or small equilibrium constant for this reaction? small, because F<sup>-</sup> is a weak base.
- Would you define F<sup>-</sup> as a proton donor or a proton acceptor? proton acceptor, because it is gaining a H<sup>+</sup> ion.
- In your written equation, is F<sup>-</sup> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? Brønsted-Lowry base, because it is gaining a proton.
- Identify the conjugate acid of F<sup>-</sup>. HF
- Identify F<sup>-</sup> as polyprotic, amphoteric, or neither. neither

6. HSO<sub>4</sub><sup>-</sup> will react with water and lose a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction.  $HSO_4^- + H_2O \rightleftharpoons SO_4^{2-} + H_3O^+$
- Would you expect a large or small equilibrium constant for this reaction? small, because HSO<sub>4</sub><sup>-</sup> is a weak acid.
- Would you define HSO<sub>4</sub><sup>-</sup> as a proton donor or a proton acceptor? proton donor, because it is losing a H<sup>+</sup> ion.
- In your written equation, is HSO<sub>4</sub><sup>-</sup> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? Brønsted-Lowry acid, because it is losing a proton.
- Identify the conjugate base of HSO<sub>4</sub><sup>-</sup>. SO<sub>4</sub><sup>2-</sup>
- Identify HSO<sub>4</sub><sup>-</sup> as polyprotic, amphoteric, or neither. amphoteric, because of both a negative charge (only will take a proton) and a H to give (can act like an acid).

# 19 3 Acids And Bases Answer Key Chemistry

**Thomas Griffiths**



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