

# PAST PAPERS

**MAY / JUNE 2014 - PAPER 12**

**0 LEVELS 4024**

# 4024 June 2014 Paper 12

**Marcel A. Müller**



**4024 June 2014 Paper 12:**

## Embracing the Track of Term: An Mental Symphony within **4024 June 2014 Paper 12**

In some sort of used by screens and the ceaseless chatter of quick transmission, the melodic beauty and mental symphony developed by the published word usually fade in to the background, eclipsed by the persistent noise and distractions that permeate our lives. However, nestled within the pages of **4024 June 2014 Paper 12** an enchanting fictional prize filled with fresh emotions, lies an immersive symphony waiting to be embraced. Constructed by an elegant musician of language, this interesting masterpiece conducts viewers on a mental journey, skillfully unraveling the hidden tunes and profound influence resonating within each carefully constructed phrase. Within the depths with this touching evaluation, we shall examine the book is central harmonies, analyze their enthralling publishing design, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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web in the context of a chemical reaction the term neutralization is used for a reaction between an acid and a base or alkali historically this reaction was represented as acid base alkali salt water for example  $\text{hcl naoh nacl h}_2\text{o}$

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web during this reaction the  $\text{h}^+$  ions from the acid and the  $\text{oh}^-$  ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web 1 sodium hydroxide solution is poured into a beaker of hydrochloric acid which contains a thermometer showing room temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable

measure an appropriate volume of each liquid eg 25 cm<sup>3</sup> 3 place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less

exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with HCl as 57.9 kJ mol<sup>-1</sup>

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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic

eg 1 HCl(aq) + NaOH(aq) → NaCl(aq) + H<sub>2</sub>O(l) for which ΔH = 57 kJ mol<sup>-1</sup> eg 2 HNO<sub>3</sub>(aq) + KOH(aq) → KNO<sub>3</sub>(aq) + H<sub>2</sub>O(l) for which ΔH = 57 kJ mol<sup>-1</sup>

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example HCl(aq) + NaOH(aq) → NaCl(aq) + H<sub>2</sub>O(aq) you can see that one mole of water is produced when one mole of HCl or one mole of NaOH reacts

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