

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HCl~~ **hydrochloric acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO₃~~ **nitric acid**
- ~~HNO₃~~ **nitric acid**
- ~~Sulfuric acid~~ **H₂SO₄**
- ~~potassium hydroxide~~ **KOH**
- ~~Sodium hydroxide~~ **NaOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~ammonium hydroxide~~ **NH₄OH**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.
HClO₄ is STRONG with a lower cell completely dissociate in water

- HCl** ~~HCl~~ ~~HNO₃~~ ~~Mg(OH)₂~~ **HNO₂** **NaOH** **NaClO₄** **NaClO₂**

3. What is the formula for hydrosulfuric acid? **H₂S**

4. MgCO_3 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for the reaction: **$\text{MgCO}_3 + \text{H}_2\text{O} \rightleftharpoons \text{Mg}^{2+} + \text{HCO}_3^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because MgCO_3 is a weak acid.**
- Would you define MgCO_3 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is MgCO_3 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of MgCO_3 : **HCO_3^-**
- Identify MgCO_3 as polyprotic, amphoteric, or neither: **polyprotic, because it contains multiple H's.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for the reaction: **$\text{F}^- + \text{H}_2\text{O} \rightleftharpoons \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F^- : **HF**
- Identify F^- as polyprotic, amphoteric, or neither: **neither.**

6. HNO_2 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for the reaction: **$\text{HNO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{NO}_2^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HNO_2 is a weak acid.**
- Would you define HNO_2 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HNO_2 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HNO_2 : **NO_2^-**
- Identify HNO_2 as polyprotic, amphoteric, or neither: **amphoteric, because of both a negative charge (only with OH^- 's neutral) and a neutral acid or the hydrogen (only not OH^- ion).**

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Emilie Sanchez



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Table of Contents 19 3 Acids And Bases Answer Key Chemistry

1. Understanding the eBook 19 3 Acids And Bases Answer Key Chemistry
 - The Rise of Digital Reading 19 3 Acids And Bases Answer Key Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Acids And Bases Answer Key Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 3 Acids And Bases Answer Key Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Acids And Bases Answer Key Chemistry
 - Personalized Recommendations
 - 19 3 Acids And Bases Answer Key Chemistry User Reviews and Ratings
 - 19 3 Acids And Bases Answer Key Chemistry and Bestseller Lists
5. Accessing 19 3 Acids And Bases Answer Key Chemistry Free and Paid eBooks
 - 19 3 Acids And Bases Answer Key Chemistry Public Domain eBooks
 - 19 3 Acids And Bases Answer Key Chemistry eBook Subscription Services
 - 19 3 Acids And Bases Answer Key Chemistry Budget-Friendly Options
6. Navigating 19 3 Acids And Bases Answer Key Chemistry eBook Formats

- ePub, PDF, MOBI, and More
 - 19 3 Acids And Bases Answer Key Chemistry Compatibility with Devices
 - 19 3 Acids And Bases Answer Key Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 3 Acids And Bases Answer Key Chemistry
 - Highlighting and Note-Taking 19 3 Acids And Bases Answer Key Chemistry
 - Interactive Elements 19 3 Acids And Bases Answer Key Chemistry
 8. Staying Engaged with 19 3 Acids And Bases Answer Key Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 3 Acids And Bases Answer Key Chemistry
 9. Balancing eBooks and Physical Books 19 3 Acids And Bases Answer Key Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 3 Acids And Bases Answer Key Chemistry
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 19 3 Acids And Bases Answer Key Chemistry
 - Setting Reading Goals 19 3 Acids And Bases Answer Key Chemistry
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 19 3 Acids And Bases Answer Key Chemistry
 - Fact-Checking eBook Content of 19 3 Acids And Bases Answer Key Chemistry
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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