

## 17.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous acids/bases.
2. Calculate hydroxide and hydronium ion concentrations and pH/pOH and the associated logarithmic functions.
3. Calculate the pH/pOH of an aqueous solution.

## Vocabulary

- |                   |                                |        |
|-------------------|--------------------------------|--------|
| 1. autoionization | 2. conjugate acid-base pair    | 3. pH  |
| 4. hydronium ion  | 5. hydronium ion concentration | 6. pOH |
|                   | 7. hydroxide ion               |        |

## Key Equations

- |   |   |
|---|---|
| 1. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 17.1) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 17.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 17.3)    | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 17.4)  |

## Part A: Completion

Use the appropriate number of significant figures in your answers. If the answer requires that you round to the nearest integer, indicate whether you are rounding up or down.

1. What products are **1** formed when hydrochloric acid ( $\text{HCl}$ ) ionizes in water? The concentration of these acids produced in 0.100 M hydrochloric acid is **0.100** M.
2. Hydrochloric acid ionizes from **1** to **2** in water. Identify **1** and **2**. Hydrochloric acid is **1** and water is **2**. The chemical formula for **1** is **HCl**.
3. **1** is the conjugate base of **2**. Identify **1** and **2**. The chemical formula for **1** is **OH<sup>-</sup>**.
4. **1** is the conjugate acid of **2**. Identify **1** and **2**. The chemical formula for **1** is **H<sub>3</sub>O<sup>+</sup>**.
5. **1** is the conjugate base of **2**. Identify **1** and **2**. The chemical formula for **1** is **OH<sup>-</sup>**.

6. **1** is the conjugate acid of **2**. Identify **1** and **2**. The chemical formula for **1** is **H<sub>3</sub>O<sup>+</sup>**.
7. **1** is the conjugate base of **2**. Identify **1** and **2**. The chemical formula for **1** is **OH<sup>-</sup>**.
8. **1** is the conjugate acid of **2**. Identify **1** and **2**. The chemical formula for **1** is **H<sub>3</sub>O<sup>+</sup>**.
9. **1** is the conjugate base of **2**. Identify **1** and **2**. The chemical formula for **1** is **OH<sup>-</sup>**.
10. **1** is the conjugate acid of **2**. Identify **1** and **2**. The chemical formula for **1** is **H<sub>3</sub>O<sup>+</sup>**.
11. **1** is the conjugate base of **2**. Identify **1** and **2**. The chemical formula for **1** is **OH<sup>-</sup>**.
12. **1** is the conjugate acid of **2**. Identify **1** and **2**. The chemical formula for **1** is **H<sub>3</sub>O<sup>+</sup>**.
13. **1** is the conjugate base of **2**. Identify **1** and **2**. The chemical formula for **1** is **OH<sup>-</sup>**.

# 19 2 Hydrogen Ions And Acidity Answer Key

**WJ Hussar**



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