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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} + 2\text{H}^{+} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{CN}^{-}$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} + 2\text{CN}^{-} + 2\text{e}^{-}$ multiplied by a factor of 2 $2\text{e}^{-} + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^{-}$ 3 the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance

being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h₂o

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reactions $6e^{-} + 14H^{+} + Cr_2O_7^{2-} + 2Cr^{3+} + 7H_2O + Fe^{2+} + Fe^{3+} + e^{-} + 2$ equalize the electrons $6e^{-} + 14H^{+} + Cr_2O_7^{2-} + 2Cr^{3+} + 7H_2O + 6Fe^{2+} + 6Fe^{3+} + 6e^{-}$ multiplied by a factor of 6

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $C_6H_8O_6 + 2H_2NO_2 + C_6H_6O_6 + 2H_2O + 2NO$ choose 1 answer $C_6H_8O_6$ and NO_2 a $C_6H_8O_6$ and NO_2 b NO_2 and $C_6H_8O_6$ c $C_6H_8O_6$ and NO_2 c h and