

No other combination of statements is used as a correct response.

MISJ2 P11

31 The gas laws can be summarised in the ideal gas equation.

$$pV = nRT$$

$$pV = \frac{m}{M_r} \cdot RT$$

where each symbol has its usual meaning.

$$p \cdot M_r = \frac{mRT}{V}$$

$$p \cdot M_r = \rho RT$$

Which statements are correct?

- 1  One mole of an ideal gas occupies the same volume under the same conditions of temperature and pressure.
- 2 The density of an ideal gas at constant pressure is inversely proportional to the temperature,  $T$ .
- 3 The volume of a given mass of an ideal gas is doubled if its temperature is raised from  $25^\circ\text{C}$  to  $50^\circ\text{C}$  at constant pressure.

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**D Keegan**

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