

Unit 8 Review

1. The balanced equation  $\text{P}_4(\text{s}) + 6\text{H}_2(\text{g}) \rightarrow 4\text{PH}_3(\text{g})$  tells us that 4.0 mol  $\text{H}_2$   
 A) reacts with 2.0 mol  $\text{P}_4$   
 B) produces 8.0 mol  $\text{PH}_3$   
 C) cannot react with phosphorus  
 D) produces 2.7 mol  $\text{PH}_3$   
 E) reacts with 4.0 mol  $\text{P}_4$

2. A 1.5-mol sample of  $\text{Fe}_2\text{O}_3$  was decomposed according to the equation:  
 $2\text{Fe}_2\text{O}_3(\text{s}) \rightarrow 2\text{Fe}_2\text{O}(\text{s}) + 3\text{O}_2(\text{g})$

How many moles of  $\text{O}_2$  are formed assuming 100% yield?

- A) 0.9 mol
- B) 1.5 mol
- C) 1.2 mol
- D) 0.7 mol
- E) 2.0 mol

3. The rusting of iron is represented by the equation  $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$ . If you have a 1.04-mol sample of iron, how many moles of  $\text{Fe}_2\text{O}_3$  will there be after the iron has rusted completely?

- A) 0.527 mol
- B) 0.790 mol
- C) 1.05 mol
- D) 1.58 mol
- E) 2.07 mol

4. Consider the following reaction, where X represents an unknown element.  
 $8\text{X}_2(\text{s}) + 2\text{B}_2\text{O}_3(\text{s}) \rightarrow \text{B}_2\text{X}_4(\text{s}) + 3\text{X}_2\text{O}_3(\text{g})$

If 175 g of X reacts completely with diboron trioxide to produce 1.43 mol of  $\text{B}_2\text{X}_4$ , what is the identity of X?

- A) Ge
- B) Mg
- C) Si
- D) Sn
- E) C

5. Refer to the following equation:  
 $4\text{NH}_3(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$

How many molecules of  $\text{NO}$  are produced when 3.02 mol of ammonia is completely reacted?

- A) 12.08
- B)  $3.84 \times 10^{23}$
- C)  $1.82 \times 10^{23}$
- D) 1.39
- E) none of these

6. In the reaction  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ , how many moles of ammonia would be produced from 1.09 mol of hydrogen and excess nitrogen?

- A) 1.42 mol
- B) 3.27 mol
- C) 0.727 mol
- D) 2.18 mol
- E) 0.363 mol

7. Refer to the following unbalanced equation:  
 $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$   
 What mass of oxygen ( $\text{O}_2$ ) is required to react completely with 13.1 g of  $\text{C}_2\text{H}_6$ ?

- A)  $2.99 \times 10^2$  g
- B) 23.1 g
- C) 4.86 g
- D) 49.2 g
- E) 0.192 g

8. How many molecules of carbon dioxide will be formed if 7.07 g of propane is burned in the following reaction?



- A)  $1.37 \times 10^{23}$  molecules
- B)  $1.03 \times 10^{23}$  molecules
- C)  $4.14 \times 10^{23}$  molecules
- D)  $3.10 \times 10^{23}$  molecules
- E)  $5.17 \times 10^{23}$  molecules

9. Calculate the mass of carbon dioxide produced from 15.7 g of octane,  $\text{C}_8\text{H}_{18}$ , in the following reaction.



- A) 79.2 g  $\text{CO}_2$
- B) 158 g  $\text{CO}_2$
- C) 89.1 g  $\text{CO}_2$
- D) 1.258 g  $\text{CO}_2$
- E) 1.400 g  $\text{CO}_2$

10. Calculate the molecules of oxygen required to react with 18.2 g of sulfur in the following reaction.



- A)  $1.03 \times 10^{23}$  molecules  $\text{O}_2$
- B)  $2.28 \times 10^{23}$  molecules  $\text{O}_2$
- C)  $1.71 \times 10^{23}$  molecules  $\text{O}_2$
- D)  $3.42 \times 10^{23}$  molecules  $\text{O}_2$
- E)  $5.13 \times 10^{23}$  molecules  $\text{O}_2$

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