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At T Answering Machines

JE Gale



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convert to basic solution $4\text{cn} + 2\text{au} + 2\text{au} + \text{cn} + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{h}_2\text{o} + \text{o}_2 + \text{h}_2\text{o}_2 + 2\text{oh}^-$ the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{ni} + \text{s} + \text{ni} + 2\text{a} + \text{q} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for

every atom of nickel oxidized but where do those electrons go

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