

# Acids & Bases HOMEWORK

## Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- HCl **hydrochloric acid**
- H<sub>2</sub>SO<sub>4</sub> **sulfuric acid**
- HNO<sub>3</sub> **nitric acid**
- HClO<sub>4</sub> **perchloric acid**
- Strontium hydroxide **LiOH**
- potassium hydroxide **KOH**
- barium hydroxide **LiOH**
- calcium hydroxide **Ca(OH)<sub>2</sub>**
- ammonium hydroxide **LiOH**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.  
*(HClO<sub>4</sub> & HNO<sub>3</sub>) both of these will completely dissociate in water*

- HCl     
  HF     
  H<sub>2</sub>S     
  Mg(OH)<sub>2</sub>     
  H<sub>2</sub>SO<sub>4</sub>     
  H<sub>2</sub>SO<sub>3</sub>     
  NH<sub>4</sub>OH     
  HClO<sub>4</sub>

3. What is the formula for hydrosulfuric acid? **H<sub>2</sub>S**

4. H<sub>2</sub>CO<sub>3</sub> will react with water and lose a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction: **H<sub>2</sub>CO<sub>3</sub> + H<sub>2</sub>O ⇌ H<sub>2</sub>O + HCO<sub>3</sub><sup>-</sup>**
- Would you expect a large or small equilibrium constant for this reaction? **small, because H<sub>2</sub>CO<sub>3</sub> is a weak acid.**
- Would you define H<sub>2</sub>CO<sub>3</sub> as a proton donor or a proton acceptor? **proton donor, because it is losing a H<sup>+</sup> ion.**
- In your written equation, is H<sub>2</sub>CO<sub>3</sub> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of H<sub>2</sub>CO<sub>3</sub>: **HCO<sub>3</sub><sup>-</sup>**
- Identify H<sub>2</sub>CO<sub>3</sub> as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's to give.**

5. F<sup>-</sup> will react with water and gain a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction: **F<sup>-</sup> + H<sub>2</sub>O ⇌ OH<sup>-</sup> + HF**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F<sup>-</sup> is a weak base.**
- Would you define F<sup>-</sup> as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H<sup>+</sup> ion.**
- In your written equation, is F<sup>-</sup> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F<sup>-</sup>: **HF**
- Identify F<sup>-</sup> as polyprotic, amphoteric, or neither. **neither.**

6. HSO<sub>4</sub><sup>-</sup> will react with water and lose a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction: **HSO<sub>4</sub><sup>-</sup> + H<sub>2</sub>O ⇌ H<sub>2</sub>O + SO<sub>4</sub><sup>2-</sup>**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO<sub>4</sub><sup>-</sup> is a weak acid.**
- Would you define HSO<sub>4</sub><sup>-</sup> as a proton donor or a proton acceptor? **proton donor, because it is losing a H<sup>+</sup> ion.**
- In your written equation, is HSO<sub>4</sub><sup>-</sup> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HSO<sub>4</sub><sup>-</sup>: **SO<sub>4</sub><sup>2-</sup>**
- Identify HSO<sub>4</sub><sup>-</sup> as polyprotic, amphoteric, or neither. **amphoteric, because of both a negative charge (only one H<sup>+</sup> to be lost) and a H to be lost in the process (one can give up acid).**

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**D Keegan**



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