

17.5

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Calculate concentrations based on the pH of a solution for hydroxide ion or hydrogen ion concentrations.
2. Calculate hydroxide ion concentrations from hydroxide ion activity and activity coefficients and calculate activity coefficients from hydroxide ion activity.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|---|-------------------|
| 1. autoionization | 2. activity coefficient (a _i) | 3. acid solution |
| 4. autoionization | 5. activity | 6. autoionization |
| | 7. autoionization | 8. pH |

Key Equations

- | | |
|---|---|
| 1. $\text{pH} = -\log_{10} a_{\text{H}^+}$ (Eq. 17.1) | 2. $\text{pOH} = -\log_{10} a_{\text{OH}^-}$ (Eq. 17.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 17.3) | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 17.4) |

Part A: Comprehension

1. Use the equilibrium constant for autoionization of water to determine the concentration of hydroxide ion in a 0.010 M hydrochloric acid solution. Assume that the autoionization of water is negligible.

2. What is the activity of H^+ in a 0.010 M hydrochloric acid solution? Assume that the activity coefficients of these species are equal to their molalities. (Eq. 17.1)
3. What is the activity of OH^- in a 0.010 M hydrochloric acid solution? Assume that the activity coefficients of these species are equal to their molalities. (Eq. 17.2)
4. What is the activity of H^+ in a 0.010 M hydrochloric acid solution? Assume that the activity coefficients of these species are equal to their molalities. (Eq. 17.3)
5. What is the activity of OH^- in a 0.010 M hydrochloric acid solution? Assume that the activity coefficients of these species are equal to their molalities. (Eq. 17.4)
6. What is the activity of H^+ in a 0.010 M hydrochloric acid solution? Assume that the activity coefficients of these species are equal to their molalities. (Eq. 17.1)
7. What is the activity of OH^- in a 0.010 M hydrochloric acid solution? Assume that the activity coefficients of these species are equal to their molalities. (Eq. 17.2)
8. What is the activity of H^+ in a 0.010 M hydrochloric acid solution? Assume that the activity coefficients of these species are equal to their molalities. (Eq. 17.3)
9. What is the activity of OH^- in a 0.010 M hydrochloric acid solution? Assume that the activity coefficients of these species are equal to their molalities. (Eq. 17.4)

10. **Autoionization**

11. **Activity**

12. **Activity**

13. **Activity**

14. **Activity**

15. **Activity**

16. **Activity**

17. **Activity**

18. **Activity**

19. **Activity**

20. **Activity**

21. **Activity**

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