

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide ion concentrations, molar ratios, and pH values for weak acid-conjugate base systems.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion | 5. molar ratio | 6. hydroxide ion |
| | 7. weak acid | 8. pH |

Key Equations

- | | |
|--|---|
| 1. $\text{pH} = -\log[\text{H}^+]$ (Eq. 14.1) | 2. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 14.3) | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair associated with the species and indicate which member of the pair is the stronger species.

1. Which is the stronger base, H_2O or OH^- ?
2. Which is the stronger acid, H_2O or H_3O^+ ?
3. Which is the stronger acid, H_2O or H^+ ?
4. Which is the stronger base, H_2O or OH^- ?
5. Which is the stronger acid, H_2O or H^+ ?
6. Which is the stronger base, H_2O or OH^- ?
7. Which is the stronger acid, H_2O or H^+ ?
8. Which is the stronger base, H_2O or OH^- ?

9. H_2O
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