

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide and hydronium ion concentrations and pH, and vice versa, for strong acids and bases.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                             |                |
|-------------------|-----------------------------|----------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. strong acid |
| 4. autoionization | 5. hydronium ion            | 6. weak acid   |
|                   | 7. weak base                | 8. pH          |

## Key Equations

- |  |   |
|--|---|
| 1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ (14.1) | 2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) |
| 3. $\text{pH} = -\log[\text{H}^+]$ (14.3)                                  | 4. $\text{pH} = \text{pK}_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right)$ (14.5)                   |

## Part A: Completion

Use the appropriate number to identify the relationship of the species indicated that are associated in the reaction. Also indicate which species acts as the conjugate acid/base.

1.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (14.2)
  - a.  $\text{H}_2\text{O}$  is the conjugate acid of  $\text{H}_2\text{O}$ .
  - b.  $\text{H}_2\text{O}$  is the conjugate base of  $\text{H}_2\text{O}$ .
  - c.  $\text{H}_3\text{O}^+$  is the conjugate acid of  $\text{H}_2\text{O}$ .
  - d.  $\text{H}_3\text{O}^+$  is the conjugate base of  $\text{H}_2\text{O}$ .
  - e.  $\text{OH}^-$  is the conjugate acid of  $\text{H}_2\text{O}$ .
  - f.  $\text{OH}^-$  is the conjugate base of  $\text{H}_2\text{O}$ .
2.  $\text{H}_2\text{O} + \text{HCl} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$  (14.2)
  - a.  $\text{H}_2\text{O}$  is the conjugate acid of  $\text{HCl}$ .
  - b.  $\text{H}_2\text{O}$  is the conjugate base of  $\text{HCl}$ .
  - c.  $\text{H}_3\text{O}^+$  is the conjugate acid of  $\text{HCl}$ .
  - d.  $\text{H}_3\text{O}^+$  is the conjugate base of  $\text{HCl}$ .
  - e.  $\text{Cl}^-$  is the conjugate acid of  $\text{HCl}$ .
  - f.  $\text{Cl}^-$  is the conjugate base of  $\text{HCl}$ .

3.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
4.  $\text{H}_2\text{O} + \text{HCl} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$
5.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
6.  $\text{H}_2\text{O} + \text{HCl} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$
7.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
8.  $\text{H}_2\text{O} + \text{HCl} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$
9.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
10.  $\text{H}_2\text{O} + \text{HCl} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$
11.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
12.  $\text{H}_2\text{O} + \text{HCl} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$
13.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
14.  $\text{H}_2\text{O} + \text{HCl} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$

# 19 2 Hydrogen Ions And Acidity Section Review Answers

**N Noddings**



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