

17.5

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations of acids and bases and vice versa.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|------------------------------|------------------------------|--------------------|
| 1. autoionization | 2. conjugate acid of a base | 3. pH |
| 4. autoionization constant | 5. conjugate base of an acid | 6. polyprotic acid |
| 7. base ionization constant | 6. hydronium ion | 7. strong acid |
| 8. base ionization constant | 7. hydronium ion | 8. strong base |
| 9. base ionization constant | 8. hydronium ion | 9. weak acid |
| 10. base ionization constant | 9. hydronium ion | 10. weak base |

Key Equations

1. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Equation 17.1)
2. $\text{pH} = -\log[\text{H}^+]$ (Equation 17.2)
3. $\text{pOH} = -\log[\text{OH}^-]$ (Equation 17.3)
4. $\text{pH} + \text{pOH} = 14.00$ (Equation 17.4)

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the reaction. Also indicate which member of the pair is the stronger acid.

1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (strong acid/strong base)
2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (strong acid/strong base)
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Barbara Janson Cohen, Kerry L. Hull



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