

# Acids & Bases HOMEWORK

## Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- HCl **hydrochloric acid**
- H<sub>2</sub>SO<sub>4</sub> **sulfuric acid**
- HNO<sub>3</sub> **nitric acid**
- HClO<sub>4</sub> **perchloric acid**
- Strontium hydroxide **LiOH**
- potassium hydroxide **KOH**
- calcium hydroxide **Ca(OH)<sub>2</sub>**
- barium hydroxide **Ca(OH)<sub>2</sub>**
- ammonium hydroxide **Li(OH)<sub>4</sub>**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.  
HClO<sub>4</sub> is STRONG with a lower pH completely dissociates in water

- HCl     
  HF     
  H<sub>2</sub>S     
  Mg(OH)<sub>2</sub>     
  HNO<sub>3</sub>     
  H<sub>2</sub>SO<sub>4</sub>     
  NH<sub>4</sub>OH     
  HClO<sub>4</sub>

3. What is the formula for hydrosulfuric acid? **H<sub>2</sub>S**

4. H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> will react with water and lose a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction. **H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> + H<sub>2</sub>O ⇌ H<sub>3</sub>C<sub>2</sub>O<sub>4</sub><sup>-</sup> + H<sub>3</sub>O<sup>+</sup>**
- Would you expect a large or small equilibrium constant for this reaction? **small, because H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is a weak acid.**
- Would you define H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> as a proton donor or a proton acceptor? **proton donor, because it is losing a H<sup>+</sup> ion.**
- In your written equation, is H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of H<sub>2</sub>C<sub>2</sub>O<sub>4</sub>. **HC<sub>2</sub>O<sub>4</sub><sup>-</sup>**
- Identify H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's to give.**

5. F<sup>-</sup> will react with water and gain a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction. **F<sup>-</sup> + H<sub>2</sub>O ⇌ HF + OH<sup>-</sup>**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F<sup>-</sup> is a weak base.**
- Would you define F<sup>-</sup> as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H<sup>+</sup> ion.**
- In your written equation, is F<sup>-</sup> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F<sup>-</sup>. **HF**
- Identify F<sup>-</sup> as polyprotic, amphoteric, or neither. **neither**

6. HSO<sub>4</sub><sup>-</sup> will react with water and lose a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction. **HSO<sub>4</sub><sup>-</sup> + H<sub>2</sub>O ⇌ SO<sub>4</sub><sup>2-</sup> + H<sub>3</sub>O<sup>+</sup>**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO<sub>4</sub><sup>-</sup> is a weak acid.**
- Would you define HSO<sub>4</sub><sup>-</sup> as a proton donor or a proton acceptor? **proton donor, because it is losing a H<sup>+</sup> ion.**
- In your written equation, is HSO<sub>4</sub><sup>-</sup> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HSO<sub>4</sub><sup>-</sup>. **SO<sub>4</sub><sup>2-</sup>**
- Identify HSO<sub>4</sub><sup>-</sup> as polyprotic, amphoteric, or neither. **amphoteric, because of both a negative charge (only will take a proton) and a H to give (can lose the proton).**

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## **Table of Contents 19 3 Acids And Bases Answer Key Chemistry**

1. Understanding the eBook 19 3 Acids And Bases Answer Key Chemistry
  - The Rise of Digital Reading 19 3 Acids And Bases Answer Key Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Acids And Bases Answer Key Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 19 3 Acids And Bases Answer Key Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Acids And Bases Answer Key Chemistry
  - Personalized Recommendations
  - 19 3 Acids And Bases Answer Key Chemistry User Reviews and Ratings
  - 19 3 Acids And Bases Answer Key Chemistry and Bestseller Lists

5. Accessing 19 3 Acids And Bases Answer Key Chemistry Free and Paid eBooks
  - 19 3 Acids And Bases Answer Key Chemistry Public Domain eBooks
  - 19 3 Acids And Bases Answer Key Chemistry eBook Subscription Services
  - 19 3 Acids And Bases Answer Key Chemistry Budget-Friendly Options
6. Navigating 19 3 Acids And Bases Answer Key Chemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - 19 3 Acids And Bases Answer Key Chemistry Compatibility with Devices
  - 19 3 Acids And Bases Answer Key Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 19 3 Acids And Bases Answer Key Chemistry
  - Highlighting and Note-Taking 19 3 Acids And Bases Answer Key Chemistry
  - Interactive Elements 19 3 Acids And Bases Answer Key Chemistry
8. Staying Engaged with 19 3 Acids And Bases Answer Key Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 19 3 Acids And Bases Answer Key Chemistry
9. Balancing eBooks and Physical Books 19 3 Acids And Bases Answer Key Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 19 3 Acids And Bases Answer Key Chemistry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine 19 3 Acids And Bases Answer Key Chemistry
  - Setting Reading Goals 19 3 Acids And Bases Answer Key Chemistry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 3 Acids And Bases Answer Key Chemistry
  - Fact-Checking eBook Content of 19 3 Acids And Bases Answer Key Chemistry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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