

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                               |                   |
|-------------------|-------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair   | 3. weak acid      |
| 4. autoionization | 5. hydronium ion ( $H_3O^+$ ) | 6. autoionization |
|                   | 7. weak base                  | 8. pH             |

## Key Equations

- |   |                         |
|---|-------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H_3O^+]$ |
| 3. $pH = -\log[H_3O^+]$                             | 4. $pOH = -\log[OH^-]$  |

## Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions that concern acid-base equilibria and solutions. Write the answers in the spaces provided.

1. What products are \_\_\_\_\_ when hydrochloric acid and sodium hydroxide react? The concentrations of these species produced in 0.100 M hydrochloric acid are \_\_\_\_\_.
2. Hydrochloric acid ionizes from \_\_\_\_\_ to \_\_\_\_\_ in water. \_\_\_\_\_ is \_\_\_\_\_ ionization. \_\_\_\_\_ is \_\_\_\_\_ ionization. \_\_\_\_\_ is \_\_\_\_\_ ionization.
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# 19 2 Hydrogen Ions And Acidity Section Review Answers

**Kevin T. Patton, Frank B. Bell, Terry  
Thompson, Peggie L. Williamson**

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