

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and an appropriate equilibrium.
2. Calculate hydroxide and hydronium ion concentrations from K_b and K_a values and the concentration of a weak acid or base.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. autoionization | 5. hydronium ion | 6. autoionization |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|--|
| 1. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ (Eq. 14.1) | 2. $K_b = \frac{[\text{B}][\text{OH}^-]}{[\text{BOH}]}$ (Eq. 14.2) |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 14.3) | 4. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions. Show calculations for all problems. Answers should include units whenever a unit is appropriate or required.

1. What products are **1** **2** formed by the autoionization of water? Write the autoionization reaction of the autoionization of water which produces an hydronium ion and a hydroxide ion.
2. Hydrochloric acid dissociates from **1** **2** in water. Write the **3** **4** autoionization reaction of the hydrochloric acid. Write the **5** **6** **7** **8** **9** **10** **11** **12** **13** **14** **15** **16** **17** **18** **19** **20** **21** **22** **23** **24** **25** **26** **27** **28** **29** **30** **31** **32** **33** **34** **35** **36** **37** **38** **39** **40** **41** **42** **43** **44** **45** **46** **47** **48** **49** **50** **51** **52** **53** **54** **55** **56** **57** **58** **59** **60** **61** **62** **63** **64** **65** **66** **67** **68** **69** **70** **71** **72** **73** **74** **75** **76** **77** **78** **79** **80** **81** **82** **83** **84** **85** **86** **87** **88** **89** **90** **91** **92** **93** **94** **95** **96** **97** **98** **99** **100** **101** **102** **103** **104** **105** **106** **107** **108** **109** **110** **111** **112** **113** **114** **115** **116** **117** **118** **119** **120** **121** **122** **123** **124** **125** **126** **127** **128** **129** **130** **131** **132** **133** **134** **135** **136** **137** **138** **139** **140** 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19 2 Hydrogen Ions And Acidity Section Review

Katrin Zwirgmaier



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