

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- HCl **hydrochloric acid**
- H₂SO₄ **sulfuric acid**
- HNO₃ **nitric acid**
- HClO₄ **perchloric acid**
- Strontium hydroxide **LiOH**
- potassium hydroxide **KOH**
- calcium hydroxide **Ca(OH)₂**
- barium hydroxide **Ca(OH)₂**
- ammonium hydroxide **Li(OH)₄**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.
HClO₄ is STRONG with a lower cell completely dissociate in water

- HCl** **HF** **H₂S** **Mg(OH)₂** **HNO₃** **Ca(OH)₂** **H₂SO₄** **H₂CO₃**

3. What is the formula for hydrosulfuric acid? **H₂S**

4. MgCO_3 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for the reaction. **$\text{MgCO}_3 + \text{H}_2\text{O} \rightleftharpoons \text{Mg}^{2+} + \text{HCO}_3^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because MgCO_3 is a weak acid.**
- Would you define MgCO_3 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is MgCO_3 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of MgCO_3 . **HCO_3^-**
- Identify MgCO_3 as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for the reaction. **$\text{F}^- + \text{H}_2\text{O} \rightleftharpoons \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F^- . **HF**
- Identify F^- as polyprotic, amphoteric, or neither. **neither**

6. HCO_3^- will react with water and lose a H^+ ion in the process.

- Write a balanced equation for the reaction. **$\text{HCO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 + \text{CO}_3^{2-}$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HCO_3^- is a weak acid.**
- Would you define HCO_3^- as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HCO_3^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HCO_3^- . **CO_3^{2-}**
- Identify HCO_3^- as polyprotic, amphoteric, or neither. **amphoteric, because of both a negative charge (only one H's) and H's (can't lose all H's in one go)**

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