

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HCl~~ **hydrochloric acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO₃~~ **nitric acid**
- ~~HNO₃~~ **nitric acid**
- ~~Sulfuric acid~~ **H₂SO₄**
- ~~potassium hydroxide~~ **KOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~potassium hydroxide~~ **KOH**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.
HClO₄ is STRONG with a lower cell completely dissociate in water

- HCl** ~~HCl~~ ~~HNO₃~~ ~~Mg(OH)₂~~ **HNO₃** **H₂SO₄** ~~HNO₂~~ **HClO₄**

3. What is the formula for hydronium? **H₃O⁺**

4. H₂CO₃ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for the reaction: **H₂CO₃ + H₂O ⇌ H₃O⁺ + HCO₃⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because H₂CO₃ is a weak acid.**
- Would you define H₂CO₃ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
- In your written equation, is H₂CO₃ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
- Identify the conjugate base of H₂CO₃: **HCO₃⁻**
- Identify H₂CO₃ as polyprotic, amphoteric, or neither: **polyprotic, because it contains multiple H's to give.**

5. F⁻ will react with water and gain a H⁺ ion in the process.

- Write a balanced equation for the reaction: **F⁻ + H₂O ⇌ HF + OH⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F⁻ is a weak base.**
- Would you define F⁻ as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H⁺ ion.**
- In your written equation, is F⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is getting a proton.**
- Identify the conjugate acid of F⁻: **HF**
- Identify F⁻ as polyprotic, amphoteric, or neither: **neither.**

6. HSO₄⁻ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for the reaction: **HSO₄⁻ + H₂O ⇌ H₃O⁺ + SO₄²⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO₄⁻ is a weak acid.**
- Would you define HSO₄⁻ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
- In your written equation, is HSO₄⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
- Identify the conjugate base of HSO₄⁻: **SO₄²⁻**
- Identify HSO₄⁻ as polyprotic, amphoteric, or neither: **amphoteric, because of both a negative charge (only one H⁺ to be lost) and a H to be lost in the process (can act like an acid).**

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