

Chimie

Acide / Base

- l'acide cède des protons, la base capte (qd une espèce peut être à la fois acide et base)
- réactions: (autoprotolyse de l'eau, réagit avec H^+ et OH^-)
- pH: potentiel hydrogène $pH = -\log\left(\frac{[H_3O^+]}{C^\circ}\right)$
- $[H_3O^+]$: ions hydronium, produit ionique de l'eau $K_w = \frac{[H_3O^+][OH^-]}{C^\circ}$ et $pH + pOH = 14$
- $pH = pK_a + \log\left(\frac{[OH^-]}{C^\circ}\right)$

absorbance

Loi de Beer-Lambert: $A = \epsilon \cdot E \cdot P \cdot C$

conductivité

conductimétrie:

Loi d'Ohm $U = R \cdot I$

conductance (G): $G = \frac{1}{R}$

conductivité (σ): $\sigma = \frac{G \cdot l}{S}$ (cm² par sec)

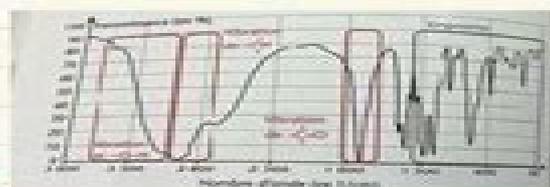
Loi de Kohlrausch: $\sigma = \sum \lambda_i \cdot |z_i| \cdot c_i$

Spectroscopie IR

Le rayonnement n'est pas suffisant pour casser les liaisons mais elle peut les faire vibrer

transmittance (en %): $T = \frac{I}{I_0}$

nombre d'onde σ : $\sigma = \frac{1}{\lambda}$



Titre massique

Une solution est composée d'un soluté et d'un solvant

$$C_m = \frac{m_{\text{soluté}}}{V_{\text{solvant}}}$$

Dilution

rajouter du solvant pour diminuer la concentra^o de soluté avec une pipette jaugée

La concentra^o ↓ mais pas la masse de soluté $m = C_{\text{mère}} \cdot V_{\text{mère}} = C_{\text{filie}} \cdot V_{\text{filie}}$

$$\text{Facteur de dilution } F = \frac{C_{\text{mère}}}{C_{\text{filie}}} = \frac{V_{\text{filie}}}{V_{\text{mère}}} = \frac{V_{\text{mère}}}{V_{\text{filie}}}$$

(on peut aussi tracer la dérivée du pH en fon^o du volume V



Chimie Ganarale Pour Inganieur

F Rizvi



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