

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations from K_b and K_a values and vice versa from K_b and K_a values.
3. Calculate K_b from K_a and vice versa.

Vocabulary

- | | | |
|-------------------|--------------------------------|--------------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. proton transfer |
| 4. hydronium ion | 5. hydronium ion concentration | 6. hydronium ion concentration |
| | 7. hydroxide ion | 8. K_b |

Key Equations

- | | |
|---|---|
| 1. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ | 2. $K_b = \frac{[\text{B}][\text{H}_3\text{O}^+]}{[\text{BH}^+]}$ |
| 3. $K_a = \frac{K_w}{K_b}$ | 4. $K_b = \frac{K_w}{K_a}$ |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the reaction. Complete the reaction.

1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (autoionization of water)
2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (autoionization of water)
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4. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
5. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
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7. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
8. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
9. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
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19 2 Hydrogen Ions And Acidity Section Review Answers

Sharyn A. Tremblay



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Table of Contents 19 2 Hydrogen Ions And Acidity Section Review Answers

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Section Review Answers
 - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Hydrogen Ions And Acidity Section Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Personalized Recommendations
 - 19 2 Hydrogen Ions And Acidity Section Review Answers User Reviews and Ratings
 - 19 2 Hydrogen Ions And Acidity Section Review Answers and Bestseller Lists
5. Accessing 19 2 Hydrogen Ions And Acidity Section Review Answers Free and Paid eBooks

- 19 2 Hydrogen Ions And Acidity Section Review Answers Public Domain eBooks
 - 19 2 Hydrogen Ions And Acidity Section Review Answers eBook Subscription Services
 - 19 2 Hydrogen Ions And Acidity Section Review Answers Budget-Friendly Options
6. Navigating 19 2 Hydrogen Ions And Acidity Section Review Answers eBook Formats
- ePub, PDF, MOBI, and More
 - 19 2 Hydrogen Ions And Acidity Section Review Answers Compatibility with Devices
 - 19 2 Hydrogen Ions And Acidity Section Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Interactive Elements 19 2 Hydrogen Ions And Acidity Section Review Answers
8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Section Review Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Section Review Answers
9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Section Review Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Section Review Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Section Review Answers
- Setting Reading Goals 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Section Review Answers
- Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Section Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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