

# Chemistry Study Guide

Periodic Table of the Elements

1	2											3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
H	He											B	C	N	O	F	Ne										
Li	Be											Al	Si	P	S	Cl	Ar										
Na	Mg	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr										
K	Ca	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe										
Rb	Sr		Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn										
Cs	Ba		Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Fl	Uup	Lv	Uus	Uuo										
Fr	Ra		Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Fl	Uup	Lv	Uus	Uuo										
		La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu											
		Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr											

## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🎯

## Matter:

Anything that has **mass** and **occupies space** 🌟🌍

Examples: air, gases, liquids, solidz!

## Atoms:

Atoms are the **smallest, indivisible particle** 🔬👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\sim 0$	-1	Orbit/Shell

# Section 2 nuclear Chemistry Study Guide Key

**John S. Phillips, Cheryl Wistrom**



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## **Table of Contents Section 2nuclear Chemistry Study Guide Key**

1. Understanding the eBook Section 2nuclear Chemistry Study Guide Key
  - The Rise of Digital Reading Section 2nuclear Chemistry Study Guide Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Section 2nuclear Chemistry Study Guide Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section 2nuclear Chemistry Study Guide Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Section 2nuclear Chemistry Study Guide Key
  - Personalized Recommendations
  - Section 2nuclear Chemistry Study Guide Key User Reviews and Ratings
  - Section 2nuclear Chemistry Study Guide Key and Bestseller Lists
5. Accessing Section 2nuclear Chemistry Study Guide Key Free and Paid eBooks

- Section 2nuclear Chemistry Study Guide Key Public Domain eBooks
  - Section 2nuclear Chemistry Study Guide Key eBook Subscription Services
  - Section 2nuclear Chemistry Study Guide Key Budget-Friendly Options
6. Navigating Section 2nuclear Chemistry Study Guide Key eBook Formats
    - ePub, PDF, MOBI, and More
    - Section 2nuclear Chemistry Study Guide Key Compatibility with Devices
    - Section 2nuclear Chemistry Study Guide Key Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Section 2nuclear Chemistry Study Guide Key
    - Highlighting and Note-Taking Section 2nuclear Chemistry Study Guide Key
    - Interactive Elements Section 2nuclear Chemistry Study Guide Key
  8. Staying Engaged with Section 2nuclear Chemistry Study Guide Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Section 2nuclear Chemistry Study Guide Key
  9. Balancing eBooks and Physical Books Section 2nuclear Chemistry Study Guide Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Section 2nuclear Chemistry Study Guide Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Section 2nuclear Chemistry Study Guide Key
    - Setting Reading Goals Section 2nuclear Chemistry Study Guide Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Section 2nuclear Chemistry Study Guide Key
    - Fact-Checking eBook Content of Section 2nuclear Chemistry Study Guide Key
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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