

Multiple Choice: (from Zumdahl *Chemistry*, 9th Edition, AP Edition)

Refer to the following diagram for questions 1 – 3

A 100.0 mL Graduated Cylinder is filled with 100.0 mL of water. A few drops of food colouring are added to the top of the water, without mixing. Initially, the top of the graduated cylinder appears to be very dark, while the bottom remains clear and colourless. After 5 min, the liquid in the graduated cylinder appears to be uniformly coloured.



1. Classify the contents in the graduated cylinder before any food colouring is added.

- (A) Element (B) Compound (C) Heterogeneous mixture
(D) Homogeneous mixture

2. Classify the contents in the graduated cylinder just after the food colouring is added.

- (A) Element (B) Compound (C) Heterogeneous Mixture (D) Homogeneous mixture

3. Classify the contents in the graduated cylinder after the 5 minutes has elapsed.

- (A) Element (B) Compound (C) Heterogeneous Mixture (D) Homogeneous mixture

4. A 10.00 g piece of metal is submerged in a graduated cylinder initially containing 20.00 mL of water. With the metal in the cylinder, the level of water is recorded as 21.40 mL. Determine the identity of the metal.

- (A) Al ($\rho = 2.70 \text{ g/mL}$)
(B) Zn ($\rho = 7.13 \text{ g/mL}$)
(C) Ag ($\rho = 10.49 \text{ g/mL}$)
(D) Au ($\rho = 19.32 \text{ g/mL}$)

5. $^{235}_{94}\text{Pu} \rightarrow X + \alpha$

In the decay represented above, what is the nucleon number of particle X?

- (A) 92 (B) 96 (C) 235 (D) 243

6. Which of the following statements about atoms is CORRECT?

- (A) Atoms are charged because they have different numbers of protons and neutrons.
(B) All atoms of an element have the same number of nucleons and electrons.
(C) An atom is mostly filled with protons.
(D) The nucleus is negatively charged.

6. When each of the following is heated to 50°C and the temperature is held constant for 5 min, which one undergoes **only** a physical change.

- (A) egg (B) steak (C) ice cream (D) cake batter

7. A sample of Ti was ionized by electron impact in a mass spectrometer. Information from the mass spectrum about the isotopes of Ti in the sample is shown in Table 2.

Table 1. Results from mass spectrum of Ti sample

m/z	46	47	48	49
Percent Abundance	9.1	7.8	74.6	8.5

Calculate the relative atomic mass of Ti in this sample. Report answers to one decimal place.

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Table of Contents Multiple Choice Review Zumdahl Ninth Edition

1. Understanding the eBook Multiple Choice Review Zumdahl Ninth Edition
 - The Rise of Digital Reading Multiple Choice Review Zumdahl Ninth Edition
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Review Zumdahl Ninth Edition
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Review Zumdahl Ninth Edition
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Review Zumdahl Ninth Edition
 - Personalized Recommendations
 - Multiple Choice Review Zumdahl Ninth Edition User Reviews and Ratings

- Multiple Choice Review Zumdahl Ninth Edition and Bestseller Lists
- 5. Accessing Multiple Choice Review Zumdahl Ninth Edition Free and Paid eBooks
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 - Multiple Choice Review Zumdahl Ninth Edition eBook Subscription Services
 - Multiple Choice Review Zumdahl Ninth Edition Budget-Friendly Options
- 6. Navigating Multiple Choice Review Zumdahl Ninth Edition eBook Formats
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 - Adjustable Fonts and Text Sizes of Multiple Choice Review Zumdahl Ninth Edition
 - Highlighting and Note-Taking Multiple Choice Review Zumdahl Ninth Edition
 - Interactive Elements Multiple Choice Review Zumdahl Ninth Edition
- 8. Staying Engaged with Multiple Choice Review Zumdahl Ninth Edition
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Choice Review Zumdahl Ninth Edition
- 9. Balancing eBooks and Physical Books Multiple Choice Review Zumdahl Ninth Edition
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Choice Review Zumdahl Ninth Edition
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Multiple Choice Review Zumdahl Ninth Edition
 - Setting Reading Goals Multiple Choice Review Zumdahl Ninth Edition
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Multiple Choice Review Zumdahl Ninth Edition
 - Fact-Checking eBook Content of Multiple Choice Review Zumdahl Ninth Edition
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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