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Emilie Sanchez



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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation

$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$

 8.3 mole $\text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} / \text{mole } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield

$$\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$$

 problems 6 & 8 are similar to problem 5

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21.3 grams of O_2 produces a 4.73% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have

$$21.3 \text{ g } \text{O}_2 \times \frac{1 \text{ mole } \text{O}_2}{32 \text{ g } \text{O}_2} = 0.666 \text{ moles } \text{O}_2$$

$$21.3 \text{ g } \text{C}_3\text{H}_6 \times \frac{1 \text{ mole } \text{C}_3\text{H}_6}{42 \text{ g } \text{C}_3\text{H}_6} = 0.507 \text{ moles } \text{C}_3\text{H}_6$$

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web a write a balanced equation for the reaction

$$2\text{Al} + 3\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$$

 b determine the theoretical yield of Al_2O_3

$$160.0 \text{ g } \text{Al} \times \frac{1 \text{ mole } \text{Al}}{27 \text{ g } \text{Al}} \times \frac{2 \text{ mole } \text{Al}_2\text{O}_3}{4 \text{ mole } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 / \text{mole } \text{Al}_2\text{O}_3 = 302.3 \text{ g } \text{Al}_2\text{O}_3$$

 c determine the percent yield

$$\frac{\text{actual yield}}{\text{theoretical yield}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$$

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