

## Acid and Base Strength Worksheet

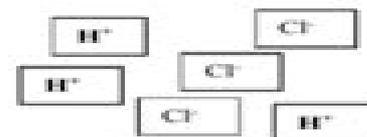
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What makes an acid *acidic* are the *excess hydrogen ions ( $H^+$ )*\* that these substances cause in a water solution.

**Strong acids, like HCl, dissociate (break up) completely** in solution, giving a maximum number of hydrogen ions. Another way to say this is, strong acids **IONIZE** completely while weak acids don't.

 HCl particles ionize completely in water to yield  $H^+$  and  $Cl^-$ .  
*Thus HCl is a STRONG ACID.*



 Only some  $H_3PO_4$  particles ionize in water. So in a solution there will be intact  $H_3PO_4$  particles as well as  $H^+$  and  $H_2PO_4^-$ . *Thus  $H_3PO_4$  is a WEAK ACID.*



So, how do you know which acids are strong and which are weak?  
**Easy; there are a few rules.**

*First, a few definitions:*

- A **binary acid** is made of **two elements**: hydrogen and something else. HCl is a binary acid made of hydrogen and chloride. Count 'em. Hydrogen (one), chloride (two). Ta-da!
- A **ternary acid** is made of **more than two elements**: hydrogen and at least two other elements.  $H_2SO_4$ , with three elements, is a ternary acid. So is  $H_3PO_4$  and acetic acid,  $HC_2H_3O_2$ . (Notice acid formulas always start with a hydrogen. These are the hydrogens that can break free in water to a lesser or greater degree depending on the strength of the acid.)

### THE RULES for deciding if an acid is strong or weak!

- **The only strong binary acids are HCl, HBr, and HI.** (Notice Cl, Br, and I are together on the periodic table.) All other binary acids are weak.
- Ternary acids are a little trickier, but not much. Compare the number of oxygens and hydrogens in the formula. **If the number of oxygens exceeds the number of hydrogens by 2 or more, the acid is strong.** All other ternary acids are weak.

$H_3PO_4$       3 hydrogens and 4 oxygens (only one more oxygen than hydrogen).    WEAK

$H_2SO_4$       2 hydrogens and 4 oxygens (two more oxygens than hydrogen).    STRONG.

\*Technically, the free  $H^+$  usually combines with a water molecule to form a hydronium ion ( $H_3O^+$ ), but we will ignore this for the time being.

# 19 2 Hydrogen Ions And Acidity Worksheet Answers

**La Motte Chemical Products  
Company, William Anthony Taylor**

## **19 2 Hydrogen Ions And Acidity Worksheet Answers:**

## Decoding **19 2 Hydrogen Ions And Acidity Worksheet Answers**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**19 2 Hydrogen Ions And Acidity Worksheet Answers**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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