

14.5

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous acids/bases.
2. Calculate hydroxide and hydronium ion concentrations and pH/pOH, and vice versa, for strong acids/bases.
3. Calculate the pH/pOH of weak acids/bases.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. hydronium ion (H_3O^+) | 6. hydroxide ion |
| | 7. weak solution | 8. pH |

Key Equations

- | | |
|---|-------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H_3O^+]$ |
| 3. $pH = -\log[H_3O^+]$ | 4. $pOH = -\log[OH^-]$ |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in (a)–(d). Label each relationship as either acid-base conjugate or neither.

- Water autoionizes to H_3O^+ and OH^- in pure water. Identify the autoionization reaction of the autoionization of water which produces an acidic solution. (1) (2) (3) (4)
- Hydrochloric acid dissociates into H^+ and Cl^- in water. Identify (1) (2) (3) (4) (5) (6) (7) (8) (9) the conjugate acid-base pair.
- Identify (1) (2) (3) (4) (5) (6) (7) (8) (9) the conjugate acid-base pair of the reaction $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$.
- Identify (1) (2) (3) (4) (5) (6) (7) (8) (9) the conjugate acid-base pair of the reaction $H_2O + HCl \rightleftharpoons H_3O^+ + Cl^-$.
- Identify (1) (2) (3) (4) (5) (6) (7) (8) (9) the conjugate acid-base pair of the reaction $H_2O + NH_3 \rightleftharpoons OH^- + NH_4^+$.
- Identify (1) (2) (3) (4) (5) (6) (7) (8) (9) the conjugate acid-base pair of the reaction $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$.

- neither
- acid-base conjugate
- acid-base conjugate
- acid-base conjugate
- neither
- acid-base conjugate
- neither

19 2 Hydrogen Ions And Acidity Answer Key

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