

Chem Honors – Acids and Bases Q and A

Define Arrhenius acid and base.

An Arrhenius acid increases the concentration of H^+ ions in water while an Arrhenius base decreases the concentration of H^+ ions in water.

What is the difference between a strong and a weak acid/base?

Strong acids/bases completely ionize in water, while weak acids/bases only partially ionize.

Explain the concept of pH.

pH is a logarithmic scale measuring the concentration of H^+ ions in a solution. A lower pH indicates higher acidity, while a higher pH indicates higher basicity.



What is the definition of conjugates?

A substance that can act as both an acid and a base.

What is the reaction between NaOH and HCl called?

This reaction is called a neutralization reaction.

List examples of strong acids and bases.

Strong acids: HCl, H_2SO_4 , HNO_3 .

Strong bases: NaOH, KOH.

What is the conjugate acid of NH_3 ?

NH_4^+ (ammonium ion)

What is the conjugate base of H_2O ?

OH^- (not shown)

Explain the Brønsted-Lowry theory of acids and bases.

Acids are proton donors, and bases are proton acceptors.

Explain the Lewis theory of acids and bases.

Acids are electron pair acceptors, and bases are electron pair donors.



LiOH
NaOH
KOH
hydroxides
of strong &
weak acids

weak acids:
 CH_3COOH
acetic acid
weak bases:
 NH_4OH

A 50.0 mL sample of nitric acid is titrated to its end point with 25.2 mL of 2.50 M KOH. What is the concentration of the acid?

mL \rightarrow L (divide by 1000)

$$M = \frac{n}{V}$$



$$V_{KNO_3} = 50 \text{ mL} = 0.050 \text{ L}$$

$$\left[\begin{array}{l} V_{KOH} = 25.2 \text{ mL} = 0.0252 \text{ L} \\ \text{Conc. KOH } M = 2.5 \text{ M} \end{array} \right]$$

$$n_{KOH} = M \cdot V$$

$$n_{KOH} = 2.5 \times 0.0252$$

$$n_{KOH} = 0.063 \text{ mol}$$

$$\rightarrow n_{KOH} = n_{HNO_3}$$

$$M = \frac{n}{V}$$

$$M_{HNO_3} = \frac{0.063}{0.050} = \underline{\underline{1.26 \text{ M}}}$$

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