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# CHEM 121 Lab Report 6 Equilibrium Reactions - Portage Learning

**Name:**

**Date:**

**Experiment: 6**

**Title: Equilibrium Reactions**

**Purpose:** the purpose of this lab is to study the equilibrium reactions by doing two experiments including the determination of the equilibrium constant  $K_c$  and the study of LeChatelier's Principle

**Procedure:**

In the small 10 ml flask that was used was filled with acetic acid and then 10 ml of propyl alcohol was added to the flask then it was mixed. Several drops of phenolphthalein indicator is added to the beaker. Then it was titrated with sodium hydroxide. Then 1 ml of it was added to a 25 ml beaker. After titration it was to become a pink color.

Next in the constant experiment we have the mixture of acetic acid and propyl alcohol and sulfuric acid catalyst. It is going to be heated for an hour so the reaction can be encouraged. The acetic acid and the propyl alcohol will react with each other to form propyl acetate + water. Which is the equilibrium. We then measure how much acetic acid is present knowing that some will be used up.

**Data/Results/Calculations:**

**Data Table:**

CH<sub>3</sub>COOH 10.00ml d=1.05g/ml MW=60

C<sub>3</sub>H<sub>7</sub>OH 10.00 ml d=0.8 g/ml MW=60

1<sup>st</sup> titration=17.1ml to check initial CH<sub>3</sub>COOH

2<sup>nd</sup> titration=17.5 ml to account for H<sub>2</sub>SO<sub>4</sub>

3<sup>rd</sup> titration= 8.2 ml to measure final CH<sub>3</sub>COOH

NaOH M= 0.508 mol/liter

Initial - CH<sub>3</sub>COOH = 10.00 ml × 1.05 g/ml / 60 g/mol × 0.020 liter = 8.75 M

Initial - C<sub>3</sub>H<sub>7</sub>OH = 10.00 ml × 0.80 g/ml / 60 g/mol × 0.020 liter = 6.67 M

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**M Mosston**



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