

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and an appropriate equilibrium.
2. Calculate hydroxide and hydronium ion concentrations from K_b and K_a values and the concentration of a weak acid or base.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. autoionization | 5. hydronium ion | 6. autoionization |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|---|--|
| 1. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ | 2. $K_b = \frac{[\text{B}][\text{OH}^-]}{[\text{BOH}]}$ |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ | 4. $\text{pH} = \log\left(\frac{1}{[\text{H}_3\text{O}^+]}\right)$ |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions. Show calculations for all problems. Answers should be written in the space provided.

1. What products are _____ when hydrochloric acid and methanol react? _____ the conjugate base of this acid is produced in this reaction. _____.
2. Hydrochloric acid ionizes in water _____ as follows: _____.
3. _____ ionizes in water to produce _____ ions. _____ ionizes _____ in water to produce _____.
4. _____ ionizes in water to form _____ and _____ ions. _____ ionizes _____ in water to produce _____.
5. _____ ionizes in water to produce _____ and _____ ions. _____ ionizes _____ in water to produce _____.

6. _____
7. _____
8. _____
9. _____
10. _____
11. _____
12. _____
13. _____
14. _____
15. _____
16. _____

19 2 Hydrogen Ions And Acidity Answer Key

**John L. Tymoczko, Jeremy M.
Berg, Lubert Stryer**



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