

Equation Sheet for A-Level Chemistry

These are the equations you need to

memorise for your Chemistry A-Level



Amount of Substance

Moles

$$n = \frac{m}{M_r}$$

Concentration

$$c = \frac{n}{V}$$

Molar gas volume

$$V = n \times 24$$

Number of particles

$$= n \times N_A$$

Ideal gas Equation

$$pV = nRT$$

Atom economy

$$= \frac{\text{M}_r \text{ of desired product}}{\text{Sum of M}_r \text{ of all reactants}} \times 100$$

% Yield

$$= \frac{\text{Actual yield}}{\text{Theoretical yield}} \times 100$$

Energetics

Enthalpy change

$$Q = mc\Delta T$$

Average bond enthalpy

$$\Delta H = \sum (\text{bond enthalpies in reactants}) - \sum (\text{bond enthalpies in products})$$

Electrode Potentials

EMF

$$E^{\ominus} = E^{\ominus} \text{right hand side} - E^{\ominus} \text{left hand side}$$

Thermodynamics

Gibbs Free Energy Equation

$$\Delta G = \Delta H - T\Delta S$$

Transition Metals

Coloured complex ions - difference in energy

$$\Delta E = h\nu = \frac{hc}{\lambda}$$

Chemical Equilibria

Equilibrium constant wrt concentration

$$K_c = \frac{[D]^d [E]^e}{[A]^a [B]^b}$$

Equilibrium constant wrt pressure

$$K_p = \frac{(D)^d (E)^e}{(A)^a (B)^b}$$

Rate Equations

$$\text{Rate} = k[A]^m[B]^n$$

Arrhenius is provided but you should be able to use:

$$k = Ae^{-E_a/RT}$$

$$\ln k = -E_a/RT + \ln A$$

Acids and Bases

$$pH = -\log_{10} [H^+]$$

$$K_w = [H^+][OH^-]$$

$$pK_a = -\log_{10} K_a$$

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

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