

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- HCl **hydrochloric acid**
- H₂SO₄ **sulfuric acid**
- HNO₃ **nitric acid**
- HClO₄ **perchloric acid**
- Strontium hydroxide **LiOH**
- potassium hydroxide **KOH**
- calcium hydroxide **CaOH**
- calcium hydroxide **Ca(OH)₂**
- ammonium hydroxide **Li(OH)₄**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.
HClO₄ is STRONG with a lower pH completely dissociates in water

- HCl HF HClO₂ Mg(OH)₂ **HNO₃** **HNO₂** H₂SO₄ **HClO₄**

- What is the formula for hydronium? **H₃O⁺**
- H₂CO₃ will react with water and lose a H⁺ ion in the process.
 - Write a balanced equation for this reaction. **H₂CO₃ + H₂O → H₃O⁺ + HCO₃⁻**
 - Would you expect a large or small equilibrium constant for this reaction? **small, because H₂CO₃ is a weak acid.**
 - Would you define H₂CO₃ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
 - In your written equation, is H₂CO₃ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
 - Identify the conjugate base of H₂CO₃. **HCO₃⁻**
 - Identify H₂CO₃ as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's to give.**
- F⁻ will react with water and gain a H⁺ ion in the process.
 - Write a balanced equation for this reaction. **F⁻ + H₂O → HF + OH⁻**
 - Would you expect a large or small equilibrium constant for this reaction? **small, because F⁻ is a weak base.**
 - Would you define F⁻ as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H⁺ ion.**
 - In your written equation, is F⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is getting a proton.**
 - Identify the conjugate acid of F⁻. **HF**
 - Identify F⁻ as polyprotic, amphoteric, or neither. **neither.**
- HCO₃⁻ will react with water and lose a H⁺ ion in the process.
 - Write a balanced equation for this reaction. **HCO₃⁻ + H₂O → H₂O⁺ + CO₃²⁻**
 - Would you expect a large or small equilibrium constant for this reaction? **small, because HCO₃⁻ is a weak acid.**
 - Would you define HCO₃⁻ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
 - In your written equation, is HCO₃⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
 - Identify the conjugate base of HCO₃⁻. **CO₃²⁻**
 - Identify HCO₃⁻ as polyprotic, amphoteric, or neither. **amphoteric, because of both a negative charge (only will take a proton) and a H to give (can act like an acid).**

19 3 Acids And Bases Answer Key Chemistry

Thomas Griffiths



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