

17.5

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Calculate concentrations based on the pH of a solution or calculate the pH of a solution based on concentration.
2. Calculate pH values and concentrations of acids and bases and the concentrations of their conjugate acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-----------------------------|--------------------------------|----------------------------------|
| 1. autoionization | 2. conjugate acid of a base | 3. pH |
| 4. autoionization constant | 5. conjugate base of an acid | 4. polyprotic acid |
| 5. base | 6. hydronium ion | 5. strong acid |
| 6. base ionization constant | 7. hydronium ion concentration | 6. weak acid |
| 7. base ionization constant | 8. hydronium ion concentration | 7. weak base |
| 8. base ionization constant | 9. hydronium ion concentration | 8. weak base ionization constant |

Key Equations

1. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Equation 17.1)
2. $[\text{H}_3\text{O}^+] = 10^{-\text{pH}}$ (Equation 17.2)
3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Equation 17.3)
4. $[\text{H}_3\text{O}^+] = 10^{-\text{pH}}$ (Equation 17.4)

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the reaction. Also indicate which member of the pair is the stronger acid.

1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (strong acid/strong base)
2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (strong acid/strong base)
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3. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
4. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
5. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
6. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
7. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
8. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
9. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$
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Victor M. Corman



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Table of Contents 19 2 Hydrogen Ions And Acidity Section Review

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Section Review
 - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Section Review
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Section Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Hydrogen Ions And Acidity Section Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Section Review
 - Personalized Recommendations
 - 19 2 Hydrogen Ions And Acidity Section Review User Reviews and Ratings

- 19 2 Hydrogen Ions And Acidity Section Review and Bestseller Lists
- 5. Accessing 19 2 Hydrogen Ions And Acidity Section Review Free and Paid eBooks
 - 19 2 Hydrogen Ions And Acidity Section Review Public Domain eBooks
 - 19 2 Hydrogen Ions And Acidity Section Review eBook Subscription Services
 - 19 2 Hydrogen Ions And Acidity Section Review Budget-Friendly Options
- 6. Navigating 19 2 Hydrogen Ions And Acidity Section Review eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 2 Hydrogen Ions And Acidity Section Review Compatibility with Devices
 - 19 2 Hydrogen Ions And Acidity Section Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Section Review
 - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Section Review
 - Interactive Elements 19 2 Hydrogen Ions And Acidity Section Review
- 8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Section Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Section Review
- 9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Section Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Section Review
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Section Review
 - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Section Review
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Section Review
 - Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Section Review
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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